



Article

An update on the mineral-like Sr-containing transition metal arsenates

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Abstract

We report on the crystal structures of three novel synthetic SrM-arsenates ($M = \text{Ni}$ and Fe^{3+}), isostructural or structurally related to the minerals from tsumcorite, carminite and brackebuschite groups. They were synthesised under mild hydrothermal conditions and further characterised using single-crystal X-ray diffraction (SXRD), scanning electron microscopy with energy dispersive spectroscopy (SEM-EDS) and Raman spectroscopy. SXRD and SEM-EDS yielded formulae: (I) $\text{SrNi}_2(\text{AsO}_4)_2 \cdot 2\text{H}_2\text{O}$, (II) $\text{Sr}_{1.4}\text{Fe}_{1.6}^{3+}(\text{AsO}_4)_2(\text{OH})_{1.6}$ and (III) $\text{SrFe}^{3+}(\text{AsO}_4)(\text{AsO}_3\text{OH})$. All three structures are built up of slightly distorted MO_6 octahedra and AsO_4 tetrahedra that are linked by Sr^{2+} with different coordination geometries and hydrogen bonds. I represent a basic structure type typical for tsumcorite-group minerals (space group $C2/m$) while II has a new intermediate structure between carminite, $\text{PbFe}_2^{3+}(\text{AsO}_4)_2(\text{OH})_2$ and arsenbrackebuschite, $\text{Pb}_2\text{Fe}^{3+}(\text{AsO}_4)_2(\text{OH})$ (s.g. Pm). III is triclinic and adopts a new structure-type (s.g. $P\bar{1}$). The structure of I is built up of infinite linear edge-sharing $\text{NiO}_4(\text{OH})_2$ octahedral chains, extending along [010] and linked by AsO_4 tetrahedra, SrO_8 polyhedra and hydrogen bonds. The structure of II is characterised by the carminite-like $\text{FeO}_4(\text{OH})_2$ octahedral chains and Edshammars-polyhedral chains, which involves SrO_{11} coordination polyhedra similar to that of PbO_{11} in arsenbrackebuschite. Both chains in II are aligned parallel to the b axis of the monoclinic unit cell and connected together by the arsenate AsO_4 tetrahedra, SrO_8 polyhedra and the hydrogen-bonding network. The compound III has a new type of crystal structure based on the unusual corrugated octahedral–tetrahedral–quadruple chains. These are made up of a central double-sided chain linked to two single-sided chains into a quadruple chain extended along the a axis. The chains in III are built up of FeO_6 octahedra and AsO_4 tetrahedra further linked to each other by shared vertices. The quadruple chains are interconnected by additional AsO_4 tetrahedra forming a heteropolyhedral 3D open framework. Strontium atoms are situated in the two channels. The structural connections to related minerals and inorganic compounds are discussed.

Keywords: metal arsenates, crystal structure, single-crystal X-ray diffraction, hydrothermal synthesis, Raman spectroscopy, environmentally relevant compounds, arsenic

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Introduction

Even though it is well known that arsenic is one of the most naturally abundant toxic metalloids, many crystal structures of arsenic oxosalts remain either undetermined or only poorly determined. This causes misunderstandings of their behaviour in various environmental media, such as soils, sediments and waters because it requires detailed knowledge of thermodynamic properties and, in particular, of formation conditions of arsenic phases and their respective crystal structures (O'Day, 2006; Drahota and Filippi, 2009). Furthermore, the crystal chemistry of transition-metal arsenates and phosphates reveals an interesting field of research for both mineralogists and solid-state researchers. Many natural or synthetic phosphates, arsenates and vanadates are of great interest due to their structural variability. In combination with the different oxidation states of some transition metals, there is a great variety of

ternary and quaternary compounds. They can foster potential applications and original physical properties such as magnetism, heterogeneous catalysis, ion exchange, optics or thermal expansion (de Pedro *et al.*, 2011; Đorđević *et al.*, 2018 and references therein). In this context, the preparation, crystal structures, spectroscopic behaviour and some physical properties of numerous alkali-, alkali-earth and transition-metal arsenic oxosalts have been investigated (Đorđević and Karanović, 2018 and references therein).

In the present contribution we are reporting on the preparation, description of the crystal structures, and spectroscopic investigation of the three novel strontium arsenates, I– $\text{SrNi}_2(\text{AsO}_4)_2 \cdot 2\text{H}_2\text{O}$, II– $\text{Sr}_{1.4}\text{Fe}_{1.6}^{3+}(\text{AsO}_4)_2(\text{OH})_{1.6}$ and III– $\text{SrFe}^{3+}(\text{AsO}_4)(\text{AsO}_3\text{OH})$. Compound I, $\text{SrNi}_2(\text{AsO}_4)_2 \cdot 2\text{H}_2\text{O}$, is isostructural with the monoclinic members of the tsumcorite-group minerals and inorganic compounds (space group $C2/m$ with $Z = 2$). Furthermore, it is the first compound synthesised and described in the $\text{SrO–NiO–As}_2\text{O}_3/\text{As}_2\text{O}_5\text{–}(\text{H}_2\text{O})$ system. Compound II, $\text{Sr}_{1.4}\text{Fe}_{1.6}^{3+}(\text{AsO}_4)_2(\text{OH})_{1.6}$, has a new crystal structure, which is an intermediate between carminite–($\text{PbFe}_2^{3+}(\text{AsO}_4)_2(\text{OH})_2$; Kharisun *et al.*, 1996) and arsenbrackebuschite-type ($\text{Pb}_2\text{Fe}^{3+}(\text{AsO}_4)_2(\text{OH})$; Hofmeister and Tillmanns, 1978) structures. Compound III, $\text{SrFe}^{3+}(\text{AsO}_4)$

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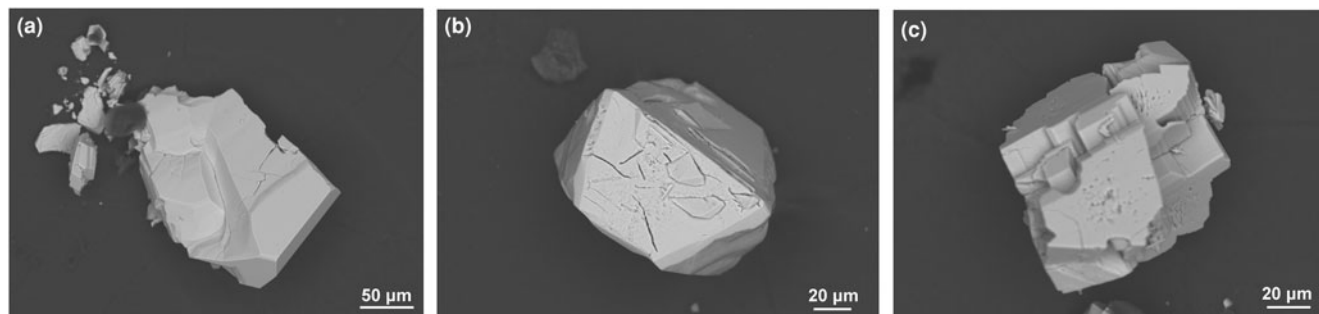


Fig. 1. Back-scattered electron images of the (a) glass-like crystal of **I**, (b) idiomorphic crystal of **II** and (c) idiomorphic crystal of **III**.

(AsO₃OH), adopts a novel structure type that is characterised by the unique octahedral–tetrahedral–quadruple [Fe(AsO₄)]_n infinite chains, which are composed of one central double- and two peripheral single-chains extended along the *a* axis. Both **II** and **III** are the first compounds synthesised in the SrO–Fe₂O₃–As₂O₃/As₂O₅–(H₂O) system.

All three SrM-arsenates adopt mixed tetrahedral–octahedral framework structures with different chains as the fundamental structural building units. The crystal-chemical features of these arsenates, their hydrogen-bonding schemes and their topological relation to the structurally-related minerals and inorganic compounds are discussed.

Experimental

Preparation of the single crystals

Compounds **I**–**III** were synthesised using low-temperature hydrothermal methods. The following reagents were used in stoichiometric quantities: Sr(NO₃)₂ (Sigma-Aldrich, 243426, 500 g), As₂O₅ (Alfa Products 87687, >99.9%), Ni(OH)₂ (Sigma-Aldrich, S58362-518, 250 g) and FeCl₂·4H₂O (Fluka 44939, 250 g).

The mixtures were transferred into polytetrafluoroethylene vessels and filled with distilled water to 80% of their inner volume. The starting mixtures had pH values of 3 for **I** and 2 for **II** and **III**. Subsequently, the mixtures were enclosed in stainless steel autoclaves and heated under autogenous pressure under the following heating and cooling conditions: the mixtures were heated from room temperature to 220°C (4 h), held at that temperature for 48 h, and cooled to room temperature (96 h). After cooling the pH values were measured as 3 for **I**, and 1 for **II** and **III**. The reaction products were filtered and washed several times using distilled water.

Compound **I** crystallised as a light green glassy-like fragment (Fig. 1) up to 0.15 mm long (yield ca. 60%) accompanied by green needles of SrNiAs₂O₇ (unpublished work) (yield ca. 40%). **II** crystallised as ruby red octahedra (yield ca. 20%) up to 120 µm long together with the colourless prisms of **III**, up to 0.2 mm long (yield ca. 80%) (Fig. 1).

Single-crystal X-ray diffraction

The intensity data for **I** and **III** were collected on a Nonius Kappa CCD single-crystal four-circle diffractometer (MoK α radiation, graphite monochromator), equipped with a 300 nm diameter capillary-optics collimator. The unit-cell parameters were determined with *HKL SCALEPACK* (Nonius, 2005–2007). A complete sphere of reciprocal space (φ and ω scans) was measured. The intensity data were processed with the Nonius program suite

DENZO-SMN (Nonius, 2005–2007) and corrected for absorption by the multi-scan method (Otwinowski *et al.*, 2003). A single-crystal of **II** was studied on a STOE StadiVari diffractometer with open Eulerian cradle, dual-beam source (sealed-tube and microfocus source) and Dectris PILATUS 300K pixel detector. Data collection for **II** was performed with X-Area (Stoe, 2013), cell refinement was done using X-Area (Stoe, 2013) and data reduction was performed with X-RED (Stoe, 2013). Single-crystal X-ray diffraction data were collected at ambient conditions, integrated and corrected for Lorentz and polarisation factors and absorption by scaling of partial multi-scans. The structures of **I**, **II** and **III** were solved by direct methods using *SHELXT* (Sheldrick, 2015a) and *WinGX* (Farrugia, 2012). The coordinates and the anisotropic displacement parameters for all non-hydrogen atoms were refined by full-matrix least-squares methods based on F^2 values by *SHELXL* (Sheldrick, 2015b).

For **I** anisotropic displacement parameters were allowed to vary for all atoms, except the hydrogen atoms. Hydrogen atoms from water molecules were located in the difference-Fourier map and refined using a riding model where $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{O})$ and with DFIX distance restraints for O–H = 0.85 Å and H...H = 1.37 Å.

For **II** the initially determined orthorhombic unit cell typical for the carminite structure type [$a = 16.500(3)$, $b = 7.5457(15)$, $c = 12.227(2)$ Å, $V = 1522.4(5)$ Å³ and $Z = 8$] gave high values for *R*-factors and weighting scheme parameters. The electron density observed was higher than desired at two of the O atom positions from the OH groups (O6 and O7), but lower at the Fe sites. Using the known model of carminite, the residual electron density could not explain such high values. Therefore, a new model was developed in space group *P1*, with the same lattice parameters. An additional symmetry check indicated that the smaller monoclinic cell [$a = 9.0719(15)$, $b = 12.227(2)$, $c = 7.5457(15)$ Å, $b = 114.575(6)$, $V = 761.2(2)$ Å³ and $Z = 4$] and the *Pm* space group is more appropriate and the unit cell was transformed using the matrix $[\frac{1}{2} \frac{1}{2} 0 / 0 \ 0 \ 1 / 0 \ \bar{1} \ 0]$. All atoms in the monoclinic unit cell are situated in general $2c$ positions excluding Sr21, Sr22, As11–As14, O21–O24, O31–O34 and mixed sites Sr61/O61–Sr64/O64, which are in special positions $1a$ and $1b$ on the mirror planes (site symmetry *m*) at $y = 0$ and $y = \frac{1}{2}$, respectively. The structure was refined as an inversion twin. The obtained Flack parameter $x = 0.570(18)$ indicate equivolume inversion twins. The transformation to monoclinic symmetry gave more than satisfactory results. The final difference-Fourier map showed no unusual features or spurious peaks around the refined positions of the atoms. The highest peak ($\Delta\rho_{\text{max}} = 0.65 \text{ e}^- \text{ Å}^{-3}$) in the map was found 0.80 Å from O11.

Table 1. Crystal data, data collection and refinement details for I–III.

	I	II	III
Crystal data			
Chemical formula	As ₂ H ₄ Ni ₂ O ₁₀ Sr	As ₂ H _{1.60} Fe _{1.60} O _{9.60} Sr _{1.40}	As ₂ HFeO ₈ Sr
<i>M_r</i>	518.91	517.08	422.32
Crystal system, space group	Monoclinic, <i>C2/m</i>	Monoclinic, <i>Pm</i>	Triclinic, <i>P1</i>
Temperature (K)	298	295	298
<i>a</i> , <i>b</i> , <i>c</i> (Å)	9.173(3), 6.2025(12), 7.562(3)	9.0719(15), 12.227(2), 7.5457(15)	6.6379(13), 9.6849(19), 12.666(3)
α , β , γ (°)	90, 114.41(2), 90	90, 114.575(6), 90	108.86(3), 90.97(3), 104.49(3)
<i>V</i> (Å ³)	391.8(2)	761.2(2)	741.8(3)
<i>Z</i>	2	4	4
Radiation type	MoK α , λ = 0.71073 Å	MoK α , λ = 0.71073 Å	MoK α , λ = 0.71073 Å
μ (mm ⁻¹)	19.97	21.45	18.01
Crystal size (mm)	0.03 × 0.02 × 0.01	0.03 × 0.02 × 0.02	0.02 × 0.02 × 0.03
Data collection			
Diffractometer	Nonius Kappa CCD	STOE StadiVari diffractometer	Nonius Kappa CCD
Absorption correction	Multi-scan	Multi-scan	Multi-scan
No. of measured, independent and observed [<i>I</i> > 2 σ (<i>I</i>)] reflections	1116, 606, 536	18131, 4785, 4001	5166, 2590, 2042
$\theta_{\max}$, $\theta_{\min}$ (°)	29.5, 4.1	30.5, 1.7	25.0, 2.4
<i>R_{int}</i>	<i>h</i> = −12→12, <i>k</i> = −8→8, <i>l</i> = −10→10	<i>h</i> = −12→12, <i>k</i> = −17→17, <i>l</i> = −10→10	<i>h</i> = −7→7, <i>k</i> = −11→11, <i>l</i> = −15→15
(<i>sin</i> σ / λ) _{max} (Å ⁻¹)	0.018	0.034	0.048
Refinement on <i>F</i>²	0.694	0.713	0.595
<i>R</i> [<i>F</i> ² > 2 σ (<i>F</i> ²)], <i>wR</i> (<i>F</i> ²), <i>S</i>	0.028, 0.060, 1.18	0.028, 0.060, 1.05	0.062, 0.141, 1.10
No. of reflections	606	4785	2590
No. of refined parameters	48	315	254
No. of restraints	2	18	1
Hydrogen site location	Difference-Fourier map	Inferred from neighbouring sites	Inferred from neighbouring sites
H-atom treatment	H atoms refined using DFIX distance restraints with O–H = 0.85 Å and <i>U</i> _{iso} (H) = 1.5 <i>U</i> _{eq} (O)	H atoms placed in geometrically calculated positions and refined using DFIX distance restraints with O–H = 0.85 Å and <i>U</i> _{iso} (H) = 1.5 <i>U</i> _{eq} (O)	H atoms placed in geometrically calculated positions and refined using the riding model (AFIX 83) with O–H = 0.82 Å and <i>U</i> _{iso} (H) = 1.5 <i>U</i> _{eq} (O)
Weighting	$w = 1/[\sigma^2(F_o^2) + (0.0186P)^2 + 3.637P]$ where $P = (F_o^2 + 2F_c^2)/3$	$w = 1/[\sigma^2(F_o^2) + (0.0178P)^2 + 1.8268P]$ where $P = (F_o^2 + 2F_c^2)/3$	$w = 1/[\sigma^2(F_o^2) + (0.0388P)^2 + 25.0224P]$ where $P = (F_o^2 + 2F_c^2)/3$
$\Delta\rho_{\max}$, $\Delta\rho_{\min}$ (e ⁻ Å ⁻³)	1.05, −0.88	0.65, −1.01	1.52, −1.24
Absolute structure	–	Refined as an inversion twin	–
Absolute structure parameter, <i>x</i>	–	0.569(18)	–
Extinction coefficient, <i>k</i>	0.0017(5) ^S		

^S*F_c** = *kF_c*[1 + 0.001 × *F_c*² λ^3 /sin(2 θ)]^{−1/4}

In order to find out how many octahedral vacancies are present, it was necessary to refine the site occupancies of four iron atoms, Fe11–Fe14, in the unit cell. The refinement of iron site occupancies shows that the Fe11 and Fe12 sites are more iron deficient than Fe13 and Fe14, which are almost fully occupied (63 and 97%, respectively). This explains the observed lower electron density at the iron position in the initially determined orthorhombic unit cell. Therefore, a simple constraint using free variables was applied: the same site occupancy was refined for Fe11 and Fe12 atoms and oxygens O61, O62 and O71 from hydroxyl groups bonding with them only, as their site occupancy factors (s.o.f.) should be equal. Similarly, the s.o.f. of oxygens O63, O64 and O72 from another hydroxyl groups should be equal to the s.o.f. of Fe13 and Fe14. The refined cation occupancies were 0.627(2) and 0.968(2) for the first F11–O61–Fe11–O71–Fe12–O62–Fe12–O71–Fe11 and second F13–O63–Fe13–O72–Fe14–O64–Fe14–O72–Fe13 array, respectively. Conclusively, oxygen sites (O61, O62 and O71) and (O63, O64 and O72) in the first and second array are occupied by oxygen atoms from OH groups in approximately 63 and 97% of the unit cells, respectively. In the remaining 37 and 3%, they are occupied by 11-coordinated strontium atoms (Sr61, Sr62 and Sr71) and (Sr63, Sr64 and Sr72), respectively, forming six mixed Sr/O sites. Because strontium

has more electrons than oxygen, the presence of the Sr atoms, even in very small quantities, make the electron density at oxygen sites greater than expected for oxygen atoms and explains the higher electron density at positions O6 and O7 in the initially determined orthorhombic unit cell. The atomic coordinates and atomic displacement parameters of the Sr and corresponding O atoms in mixed Sr/O sites were refined as identical.

In the latter cycles of refinement the presence of disordered Sr71 atoms was found, resulting in the appearance of two additional partly occupied Sr sites of very low occupancy (ca. 10% of available sites), denoted as Sr81 and Sr82 and separated from Sr71 by a distance of only 0.728(6) and 0.697(5) Å, respectively. Similarly, the Sr72 position is also partly occupied and disordered, resulting in additional Sr83 and Sr84 atoms of very low occupancy (ca. 10% of available sites) at distances of 0.918(8) and 0.891(9) Å, respectively. This finding suggests a statistical disorder and movement of Sr71 and Sr72 atoms between these Sr81–Sr82 and Sr83–Sr84 neighbouring sites, respectively. These movements are directed to the neighbouring iron-site vacancies. Hydrogen atoms of the OH groups (OH61, OH62, OH63, OH64, OH71 and OH72) were placed in geometrically calculated positions and refined using a riding model where *U*_{iso}(H) = 1.5*U*_{eq}(O) and DFIX distance restraints for O–H = 0.85 Å and H...A = 1.95 Å.

Table 2. Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ($\text{\AA}^2$) for **I**.

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Sr1	0	1/2	0	0.0205(2)
As1	−0.41918(7)	1/2	−0.22558(9)	0.01074(18)
Ni1	1/4	3/4	1/2	0.0131(2)
O1	0.3267(5)	1	0.3814(8)	0.0230(11)
O2	0.3056(5)	1/2	0.3556(6)	0.0163(9)
O3	−0.0310(3)	0.7745(5)	−0.2685(4)	0.0176(6)
O4	−0.2835(6)	1/2	0.0045(7)	0.0304(12)
H1	0.413(2)	1	0.365(3)	0.034*
H2	0.253(3)	1	0.265(4)	0.034*

For **III** all non-hydrogen atoms excluding Sr2, which displays an extremely high degree of positional disorder, were initially refined anisotropically. The refinement in the lowest symmetry space group *P1* did not provide an ordered structure. However, in space group *P1* two disordered Sr2 atoms in the unit cell appear distributed among 12 positions. Two pairs of sites were related by the pseudo-inversion centre, but the remaining eight were not. This means that Sr2 is distributed among at least 10 positions in *P1*, i.e. on 10 crystallographically independent sites Sr21–Sr30 with a multiplicity of two. Because Sr21–Sr30 atoms are spatially close they were refined with equal atomic displacement parameters. Their occupancies were refined using the free variables and SUMP instruction. The refined positions showed site occupancy factors from 0.225(6) to 0.041(4). Owing to the disorder of the Sr2 site, *R* parameters are large ($R[F^2 > 2\sigma(F^2)] = 0.06$ $wR(F^2) = 0.14$). However, it can be assumed that the structure determined is correct, except possibly for details.

Hydrogen atoms from OH groups were placed in geometrically calculated positions and refined using the riding model with O–H = 0.82 Å and isotropic displacement parameters $U_{\text{iso}}(\text{H})$ fixed to 1.5 times the $U_{\text{eq}}(\text{O})$ value of the atoms they are linked to.

The O34–H34...H44–O44 contacts are short to host unit-occupancy H atoms. As a consequence, O34 and O44 atoms serve as both the hydrogen bond donor and acceptor, to each other, alternatively. The donor–acceptor O34...O44 distance is 2.832(17) Å and the H34...H44 distance is 1.4827(5) Å. Therefore the H34 and H44 atom position occupancy is a maximum of 50%, which is consistent with the charge balance.

Relevant information on crystal data, data collection and refinements are compiled in Table 1. Fractional atomic coordinates and equivalent isotropic displacement parameters for **I–III** are presented in Tables 2–4 and anisotropic displacement parameters for **I–III** are presented in Supplementary Tables S1–S3. Selected bond lengths for non-H atoms are listed in Table 5 and Tables S4–S6. Bond valence analysis obtained using VALIST software (Wills, 2010) are presented in Tables S7–S9. Interatomic distances and angles involving H atoms are given in Tables 6–8. All drawings were plotted using ATOMS (Dowty, 2000) and VESTA (Momma and Izumi, 2011). The crystallographic information files and Supplementary Tables S1–S9 have been deposited with the Principal Editor of *Mineralogical Magazine* and are available as Supplementary material (see below).

Elemental Analysis

Elemental analysis was performed on the hydrothermally synthesised crystals with a JEOL JSM-6610 LV scanning electron microscope (SEM) connected to an Oxford INCA Energy 350 EDX

Table 3. Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ($\text{\AA}^2$) for **II**.

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
Sr11	0.9787(9)	0.24995(5)	0.9845(10)	0.01249(12)	
Sr21	0.4844(9)	0.000000	0.9870(11)	0.01321(17)	
Sr22	0.4851(9)	0.500000	0.4879(11)	0.01309(17)	
As11	0.8930(9)	0.000000	1.1968(10)	0.0069(3)	
As12	0.8934(9)	0.500000	0.6871(10)	0.0070(3)	
As13	1.0660(9)	0.000000	0.7647(10)	0.0075(3)	
As14	1.0663(9)	0.500000	1.2922(10)	0.0073(3)	
As21	1.4036(9)	0.24996(8)	1.1971(10)	0.0063(2)	
As22	0.5620(9)	0.25007(8)	0.7763(10)	0.0071(2)	
Fe11	0.7333(9)	0.3668(2)	1.2326(11)	0.0117(5)	0.628(2)
Fe12	0.7328(9)	0.1331(2)	0.4921(11)	0.0114(5)	0.628(2)
Fe13	1.2287(9)	0.36708(12)	0.7451(10)	0.0081(3)	0.968(2)
Fe14	1.2285(9)	0.13280(12)	1.4750(10)	0.0082(3)	0.968(2)
O61	0.6444(10)	0.500000	1.0698(11)	0.0233(8)	0.628(2)
O62	0.6440(10)	0.000000	0.5658(11)	0.0222(8)	0.628(2)
O71	0.8098(10)	0.2501(3)	1.4004(11)	0.0220(8)	0.628(2)
Sr61	0.6444(10)	0.500000	1.0698(11)	0.0233(8)	0.372(2)
Sr62	0.6440(10)	0.000000	0.5658(11)	0.0222(8)	0.372(2)
Sr71	0.8098(10)	0.2501(3)	1.4004(11)	0.0220(8)	0.2975(18)
Sr81	0.782200	0.199306	1.426899	0.0220(8)	0.0372(2)
Sr82	0.784300	0.298903	1.348999	0.0220(8)	0.0372(2)
O63	1.3140(14)	0.500000	0.9180(16)	0.018(2)	0.968(2)
O64	1.3128(14)	0.000000	1.3858(16)	0.024(2)	0.968(2)
O72	1.1365(12)	0.2492(6)	0.5634(16)	0.0205(18)	0.968(2)
Sr63	1.3140(14)	0.500000	0.9180(16)	0.018(2)	0.032(2)
Sr64	1.3128(14)	0.000000	1.3858(16)	0.024(2)	0.032(2)
Sr72	1.1365(12)	0.2492(6)	0.5634(16)	0.0205(18)	0.0259(18)
Sr83	1.211700	0.298000	0.643209	0.0205(18)	0.0032(2)
Sr84	1.213600	0.199903	0.580900	0.0205(18)	0.0032(2)
O11	0.9471(12)	0.3882(6)	1.2194(14)	0.0124(16)	
O12	0.9479(12)	0.1110(6)	0.7184(14)	0.0115(15)	
O13	1.0107(13)	0.3873(6)	0.7554(14)	0.0158(17)	
O14	1.0096(12)	0.1135(6)	1.2458(14)	0.0150(17)	
O21	0.7852(16)	0.500000	0.4435(18)	0.020(2)	
O22	0.7875(15)	0.000000	1.3346(17)	0.016(2)	
O23	1.1667(14)	0.500000	1.5445(16)	0.0074(18)	
O24	1.1673(15)	0.000000	0.6128(17)	0.012(2)	
O31	0.7609(15)	0.000000	0.9653(16)	0.020(2)	
O32	0.7592(15)	0.500000	0.7858(17)	0.022(2)	
O33	1.2091(15)	0.000000	0.9916(16)	0.020(2)	
O34	1.2081(16)	0.500000	1.2076(17)	0.020(2)	
O41	0.6824(13)	0.2515(7)	1.0143(14)	0.0223(18)	
O42	0.6819(13)	0.2437(6)	0.6590(14)	0.0202(18)	
O43	1.2778(12)	0.2606(6)	0.9602(14)	0.0106(14)	
O44	1.2778(12)	0.2444(6)	1.3093(14)	0.0130(15)	
O51	1.5219(12)	0.3629(7)	1.2646(14)	0.0149(16)	
O52	1.5224(12)	0.1393(6)	1.2489(14)	0.0139(16)	
O53	0.4418(12)	0.3612(6)	0.7191(14)	0.0119(15)	
O54	0.4415(12)	0.1375(6)	0.7139(14)	0.0135(16)	
H61	0.649(6)	0.500000	0.959(5)	0.035*	0.628(2)
H62	0.649(6)	0.000000	0.681(2)	0.033*	0.628(2)
H63	1.251(5)	0.500000	0.976(5)	0.027*	0.968(2)
H64	1.248(5)	0.000000	1.266(2)	0.036*	0.968(2)
H71	0.9037(18)	0.259(14)	1.492(9)	0.033*	0.628(2)
H72	1.0427(17)	0.229(6)	0.546(11)	0.031*	0.968(2)

analysis unit. The single crystals were mounted with conducting carbon onto Al sample holders. The sample holder was then coated with a thin layer of carbon. An accelerating voltage of 20 kV and accumulation time of 20 s were used for the chemical analysis. As a qualitative measure, energy dispersive spectroscopy confirmed the presence of each reported element in the title compounds. Semi-quantitative chemical analyses gave the presence of strontium in the range of 17.0–17.4 at.%, arsenic in the range 28.8–29.3 at% and nickel in the range 17.8–18.9 at% for **I**, strontium in the range of 26.8–27.4 at.%, arsenic in the range 28.3–28.4 at.% and iron in the range 12.9–15.2 at.% for **II** and

Table 4. Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ($\text{\AA}^2$) for **III**.

	<i>x</i>	<i>y</i>	<i>z</i>	<i>U</i> _{iso} */ <i>U</i> _{eq}	Occ. (<1)
Sr1	0.3740(2)	0.78866(17)	0.55533(12)	0.0267(4)	
Sr21	−0.470(2)	0.3620(13)	0.9942(8)	0.0397(15)	0.210(6)
Sr22	−0.539(2)	0.4096(12)	0.9889(7)	0.0397(15)	0.225(6)
Sr23	−0.015(2)	0.5855(15)	0.7910(15)	0.0397(15)	0.140(6)
Sr24	0.004(3)	0.5890(19)	0.7438(19)	0.0397(15)	0.113(6)
Sr25	0.345(9)	0.412(4)	1.000(3)	0.0397(15)	0.054(5)
Sr26	0.205(6)	0.382(3)	0.9910(19)	0.0397(15)	0.072(4)
Sr27	0.363(8)	0.646(4)	0.978(3)	0.0397(15)	0.052(5)
Sr28	0.257(7)	0.687(5)	0.951(3)	0.0397(15)	0.043(4)
Sr29	−0.086(6)	0.656(4)	0.995(2)	0.0397(15)	0.055(4)
Sr30	0.119(7)	0.627(5)	0.890(3)	0.0397(15)	0.041(4)
As1	1.2429(2)	0.98981(17)	0.87033(11)	0.0186(3)	
As2	0.7519(2)	1.18115(17)	0.74450(11)	0.0212(4)	
As3	0.8813(2)	0.81547(17)	0.57294(11)	0.0186(3)	
As4	0.4843(2)	0.51377(17)	0.70512(12)	0.0244(4)	
Fe1	0.6981(3)	0.8893(2)	0.83659(16)	0.0199(5)	
Fe2	0.2395(3)	1.1634(2)	0.68151(15)	0.0189(5)	
O11	0.9993(14)	0.8829(12)	0.8546(8)	0.026(2)	
O12	1.3836(15)	0.8782(12)	0.7971(8)	0.025(2)	
O13	1.3401(15)	1.0702(13)	1.0069(8)	0.031(3)	
O14	1.2603(16)	1.1421(11)	0.8287(8)	0.025(2)	
O21	0.7576(17)	1.1040(11)	0.8434(8)	0.028(2)	
O22	0.9504(15)	1.1796(13)	0.6659(8)	0.030(3)	
O23	0.5313(14)	1.1118(12)	0.6537(7)	0.022(2)	
O24	0.7743(18)	1.3656(12)	0.8252(8)	0.037(3)	
H24	0.7732	1.4161	0.7844	0.056*	
O31	1.1159(14)	0.9371(10)	0.6108(7)	0.020(2)	
O32	0.7108(16)	0.8241(13)	0.6695(7)	0.033(3)	
O33	0.7247(15)	0.8141(12)	0.4671(8)	0.025(2)	
O34	0.9536(16)	0.6487(12)	0.5356(10)	0.036(3)	
H34	0.8488	0.5767	0.5144	0.055*	0.5
O41	0.6038(15)	0.6643(11)	0.8139(8)	0.023(2)	
O42	0.3585(17)	0.3820(11)	0.7581(8)	0.031(3)	
O43	0.3137(17)	0.5583(13)	0.6294(9)	0.037(3)	
O44	0.6633(17)	0.4507(12)	0.6187(8)	0.034(3)	
H44	0.7087	0.5105	0.5858	0.051*	0.5

strontium of 15.4 at.%, arsenic of 34.5 at.% and iron of 10.3 at.% for **III**. The atomic proportions of constituent elements calculated from atomic % are Sr:Ni:As = 1.2:1.3:2.0 for **I**, Sr:Fe:As = 1:1.9:2 for **II** and Sr:Fe:As = 0.9:1.1:2 for **III** and correspond quite well with the ratios obtained from the structural analyses.

Raman spectroscopy

To obtain further information on the anion groups, Raman spectra were acquired. Non-oriented single-crystal Raman spectra of **I–III** on the randomly oriented single-crystals were collected using a Horiba LabRam HR Evolution system equipped with a Si-based, Peltier-cooled CCD detector in the spectral range from 4000 to 60 cm^{-1} . The 633 nm excitation line of a He–Ne laser was focused with a 50× objective on the randomly oriented single crystals of **I–III**. The samples spectra were acquired with a nominal exposure time of 10 s (confocal mode, Olympus 1800 lines/mm, 1.5 μm lateral resolutions and $\sim 3 \mu\text{m}$ depth resolution).

Results and discussion

Description of the crystal structures

Crystal structure of $\text{SrNi}_2(\text{AsO}_4)_2 \cdot 2\text{H}_2\text{O}$ (**I**)

The basic building units in the structure of **I** are infinite linear chains of edge-sharing $\text{NiO}_4(\text{OH}_2)_2$ octahedra, extending along

Table 5. Selected bond distances ($\text{\AA}$) in the coordination polyhedra of the compounds **I–III**.

	I	II	II	III
M–O	2.055(3)	1.841(4)	1.924(8)	1.926(9)
M–O	2.055(3)	2.000(7)	1.982(6)	1.990(10)
M–O	2.054(3)	1.997(4)	2.023(6)	2.013(9)
M–O	2.054(3)	2.031(7)	2.025(6)	2.027(9)
M–O	2.078(3)	2.069(8)	2.027(7)	2.033(10)
M–O	2.078(3)	2.187(7)	2.131(5)	2.107(10)
<M–O>	2.0623	2.0208	2.0187	2.0160
M–O		1.847(4)	1.906(8)	1.948(9)
M–O		1.995(4)	2.022(7)	1.964(11)
M–O		2.006(7)	2.033(7)	1.975(9)
M–O		2.026(7)	2.025(6)	1.976(10)
M–O		2.027(7)	2.022(7)	2.022(9)
M–O		2.190(7)	2.123(6)	2.126(9)
<M–O>		2.0152	2.0218	2.0018

[010] (Fig. 2a). The coordination environment of Ni1 generates a tetragonal slightly elongated octahedron $\text{NiO}_4(\text{OH}_2)_2$ having two pairs of symmetry equivalent O atoms (two O1 at a distance of 2.055(3) and two O3 at a distance of 2.054(3), Table S4) and two apical O2 at a distance of 2.078(3) $\text{\AA}$. The $\text{NiO}_4(\text{OH}_2)_2$ octahedra are almost perfect with an average <Ni1–O> distance of 2.062 $\text{\AA}$, which is in agreement with the value of 2.09 $\text{\AA}$ calculated from the sum (0.69 + 1.4 = 2.09 $\text{\AA}$) of effective ionic radii (Shannon, 1976).

Adjacent octahedral chains are linked to each other by sharing corners with AsO_4 tetrahedra (Fig. 2a, b). The individual As–O bond lengths vary from 1.669(5) to 1.702(4) $\text{\AA}$, the average <As–O> distance is 1.685 $\text{\AA}$ which is close to the sum of the effective ionic radii of As^{+5} in tetrahedral coordination and O^{2-} (0.335 + 1.4 = 1.735 $\text{\AA}$; Shannon, 1976). The linear octahedral chains and AsO_4 tetrahedra are arranged in layers parallel to (001) (Fig. 2a, b). The topology of the layer is identical to that in natrochalcite, $\text{NaCu}_2(\text{SO}_4)_2(\text{H}_3\text{O}_2)$ (Krickl *et al.*, 2007; Hawthorne *et al.*, 2000; Hawthorne, 2014).

Between these layers are situated 8-coordinated (6+2) Sr^{2+} cations. The mutually isolated SrO_8 coordination polyhedra can be described as a somewhat distorted orthorhombic prism or dicapped slightly elongated octahedron, where the four symmetry equivalent O3 oxygen atoms are located in the equatorial plane and two O4 are at the vertices. In addition, two opposite triangular faces are capped by two O2 atoms at longer distances. The Sr–O distances are in the range 2.573(3)–2.976(5) $\text{\AA}$ and the average value is 2.684 $\text{\AA}$ (Table S4), which is in agreement with the value of 2.66 $\text{\AA}$ calculated from the sum (1.26+1.4 = 2.66 $\text{\AA}$) of effective ionic radii for Sr^{2+} with coordination number (CN) = 8 (Shannon, 1976).

The other way of describing the crystals structure of **I** are two kind of slabs positioned alternately along the [001] direction (Fig. 2b). The SrO_8 polyhedra share two edges and four vertices with six adjacent AsO_4 tetrahedra forming the first slab situated between *z* at approximately −0.33 and 0.33. The thickness of this slab is approximately double the thickness of the second slab which hosts $\text{NiO}_4(\text{OH}_2)_2$ octahedra chains combined with intralayer hydrogen bonds. Its boundaries lie between *z* at approximately 0.33 and 0.67.

Bond-valence calculations (Brown and Altermatt, 1985; Brese and O’Keeffe, 1991) show that the Sr–O, Ni–O and As–O bond lengths are consistent with the presence of Sr^{2+} , Ni^{2+} , As^{5+} and

Table 6. Hydrogen-bond geometry (Å, °) for **I**.

D–H...A	D–H	H...A	D...A	D–H...A
O1–H1...O1 ^{xv}	0.85(2)	2.36(2)	2.935(10)	125(1)
O1–H1...O3 ^{xvi}	0.85(2)	2.29(1)	2.909(5)	130(1)
O1–H1...O3 ^{xvii}	0.85(2)	2.29(1)	2.909(5)	130(1)
O1–H2...O4 ^{xiii}	0.86(2)	1.86(4)	2.601(7)	144(4)

Symmetry codes: (xiii) $x+1/2, y+1/2, z$ (xv) $-x+1, -y+2, -z+1$; (xvi) $-x+1/2, -y+3/2, -z$; (xvii) $-x+1/2, y+1/2, -z$.

Table 7. Hydrogen-bond geometry (Å, °) for **II**.

D–H...A	D–H	H...A	D...A	D–H...A
O61–H61...O32	0.85(2)	1.95(2)	2.745(8)	155(5)
O62–H62...O31	0.86(2)	1.95(2)	2.748(8)	154(5)
O63–H63...O34	0.85(2)	1.95(2)	2.728(10)	153(5)
O64–H64...O33	0.85(2)	1.95(2)	2.721(10)	151(4)
O71–H71...O72 ^x	0.84(2)	1.95(2)	2.695(5)	147(3)
O72–H72...O71 ^{vii}	0.85(2)	1.95(2)	2.695(5)	145(3)

Symmetry code: (vii) $x, y, z-1$; (x) $x, y, z+1$.

O²⁻ ($\Sigma v_{ij} = 1.886(6)$ valence units (vu) for Sr1 with CN = 8, $\Sigma v_{ij} = 2.106(6)$ vu for Ni1 with CN = 6, $\Sigma v_{ij} = 4.991(1)$ vu for As1 with CN = 4). If four symmetry equivalent O4 oxygen atoms at very long distances of 3.675(3) Å are included, the bond valence sum Σv_{ij} for Sr1 is 1.944 vu (Table S7). The bond valence sum for 4-coordinated O2, which is not included in hydrogen bonding, is close to the nominal valence of 2- ($\Sigma v_{ij} = 1.962(2)$ vu). Without taking into account the contribution from hydrogen atoms, the sum of the bond valences around another three O atoms in the structure is lower: $\Sigma v_{ij} = 0.72(6)$, 1.90(5) and 1.56(2) vu which contribute less than 2% each to the bond valence sum), respectively. This indicates that atoms O1 and O4 are preferentially involved in the hydrogen bonding.

The occupancy of the *M* site with divalent Ni²⁺ in **I** causes a complete occupation of the O1 sites by H₂O molecules and the occurrence of one interlayer and one intralayer hydrogen bond. The strong interlayer hydrogen bonds are placed across the layers between O1 and O4 (Fig. 2a–c). The weak intralayer hydrogen bonds extend between two symmetry equivalent O1 atoms belonging to two adjacent octahedral chains (Table 6). The O1–O4^{xiii} distance of 2.601(7) Å is rather short, while another donor–acceptor distance O1–O1^{ix} of 2.935(10) Å is in the usual range for common hydrogen bonds. The first is straight and strong, while the second is a bent and weak hydrogen bond (O1–H2...O4^{xiii} and O1–H1...O1^{xv} angles are 144(4) and 125(1)°, respectively; symmetry codes: (xv) $-x+1, -y+2, -z+1$; (xiii) $x+1/2, y+1/2, z$).

Unlike **I** and its phosphate analogue SrNi₂(PO₄)₂·2H₂O (Assani *et al.*, 2010), in the majority of tsumcorite-group members intralayer hydrogen bonds are usually much shorter and stronger than the interlayer hydrogen bonds. According to Effenberger *et al.* (2000) an increase in the O1–O1 distance indicates the absence of strong symmetry restricted hydrogen bonds between two symmetry equivalent O1 atoms. In the structure of the **I** hydrogen atom H1 is moved away from the inversion centre and therefore the intermolecular O1–H1...H1–O1 contacts became long enough to host unit-occupancy H1 atoms. As a consequence, two symmetry equivalent O3 atoms became suitable to act as acceptors of adjacent H1 (donor–acceptor distance is 2.909(5) Å for O1–H1...O3^{xvi} and O1–H1...O3^{xvii}; symmetry

Table 8. Hydrogen-bond geometry (Å, °) for **III**.

D–H...A	D–H	H...A	D...A	D–H...A
O24–H24...O44 ^{xi}	0.82	2.36	3.114(16)	153
O34–H34...O43 ⁱⁱ	0.82	1.97	2.631(16)	137
O34–H34...O44	0.82	2.26	2.833(15)	128
O44–H44...O34	0.82	2.08	2.833(15)	152
O44–H44...O43 ⁱⁱ	0.82	2.58	3.123(14)	125

Symmetry codes: (ii) $-x+1, -y+1, -z+1$; (xi) $x, y+1, z$.

codes: (xvi) $-x+1/2, -y+3/2, -z$; (xvii) $-x+1/2, y+1/2, -z$). These bonds are weak and bent (O1–H1...O3^{xvi} and O1–H1...O3^{xvii} angles are 130(1)°). In such a way O1 is involved in forming weak trifurcated hydrogen bonds (O1–H1...O1^{xv}, O3^{xvi}, O3^{xvii}) that are placed between chains in one layer and strong O1–H2...O4^{xiii} interlayer hydrogen bonds located between layers.

Crystal structure of Sr_{1.4}Fe_{1.6}(AsO₄)₂(OH)_{1.6} (**III**)

Compound **II** is only partly (ca. 80%) isostructural with carminite, Pb²⁺Fe₂³⁺(AsO₄)₂(OH)₂, which means that there are regions where the carminite-type structure is maintained in the form of octahedral Fe(O,OH)₆ chains (Fig. 3a). In contrast to carminite, the structure of **II** contains four partially vacant Fe and six mixed Sr/O sites in which 11-coordinated Sr²⁺ and oxygen atoms from OH groups statistically share the same position. Therefore, in addition to edge-sharing octahedral Fe₂(O,OH)₁₀ dimers that build infinite octahedral chains running along [010], such as in carminite, a new type of chain is formed. These new chains (Figs 3b, c) are similar to the PbO₁₁-chains in arsenbrackebuschite and can be described as trapezoidal face-sharing Sr centred Edshammar polyhedra connected along [010].

The parallel octahedral Fe(O,OH)₆-chains and Edshammar polyhedral SrO₁₁-chains are distributed statistically over two parallel rows in the unit cell. The octahedral Fe(O,OH)₆-chain in the first row consists of edge-shared Fe11₂(O,OH)₁₀ and Fe12₂(O,OH)₁₀ octahedral dimers which are linked in chains sharing corners. In the second row, the chain is formed by similar edge-shared Fe13₂(O,OH)₁₀ and Fe14₂(O,OH)₁₀ octahedral dimers (Fig. 3a). All octahedral dimers share an edge along mirror planes at $y=0$ and 0.5 in the first and second row, respectively. According to the refinement results, the volume occupied by the Fe11₂(O,OH)₁₀–Fe12₂(O,OH)₁₀ chain is 63% and then by the Fe13₂(O,OH)₁₀–Fe14₂(O,OH)₁₀ chain is 97%. This means that the Edshammar polyhedral SrO₁₁-chains occupy the first and second row by 37 and 3%, respectively.

Four iron cations Fe11–Fe14, situated at general positions on the pseudoinversion centre, are coordinated by the four oxygen anions belonging to the AsO₄ tetrahedra and two oxygen anions belonging to the hydroxyl groups. The Fe(O,OH)₆ dimeric octahedra are distorted showing one shorter, four typical and one longer Fe–O bonds close to the calculated value from effective ionic radii of 6-coordinated Fe³⁺ cations (0.65 + 1.4 = 2.05 Å; Shannon, 1976). Oxygen from the common edge of the dimer and the oxygen from the hydroxyl group bridging two adjacent dimers in the octahedral chain form the longest and the shortest Fe–O interatomic distances, respectively. The average <Fe³⁺–O> bond distances are similar and very close to the calculated value: 2.021, 2.015, 2.019 and 2.022 Å for Fe11(O,OH)₆, Fe12(O,OH)₆, Fe13(O,OH)₆ and Fe14(O,OH)₆ octahedra, respectively (Table 5).

All arsenic atoms are coordinated by four nearest oxygens, adopting a typical tetrahedral coordination. In the carminite

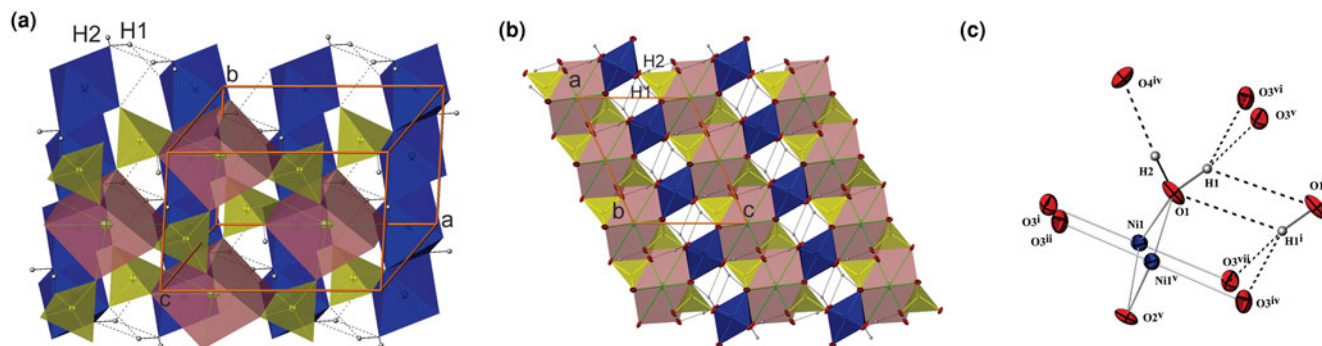


Fig. 2. (a) The edge-sharing $\text{NiO}_4(\text{OH})_2$ -octahedra chains (blue) extending along $[010]$ and bonded with As_{10}_4 -tetrahedra in a layer parallel to (001) which are linked by dicated Sr_{10}_8 octahedra (dark pink) in **I**; (b) the projection of **I** along the b axis showing two types of slabs parallel to (001) ; (c) the interlayer and intra-layer hydrogen bonds in **I**. H atoms are presented as small spheres and H-bonds are shown as the broken lines.

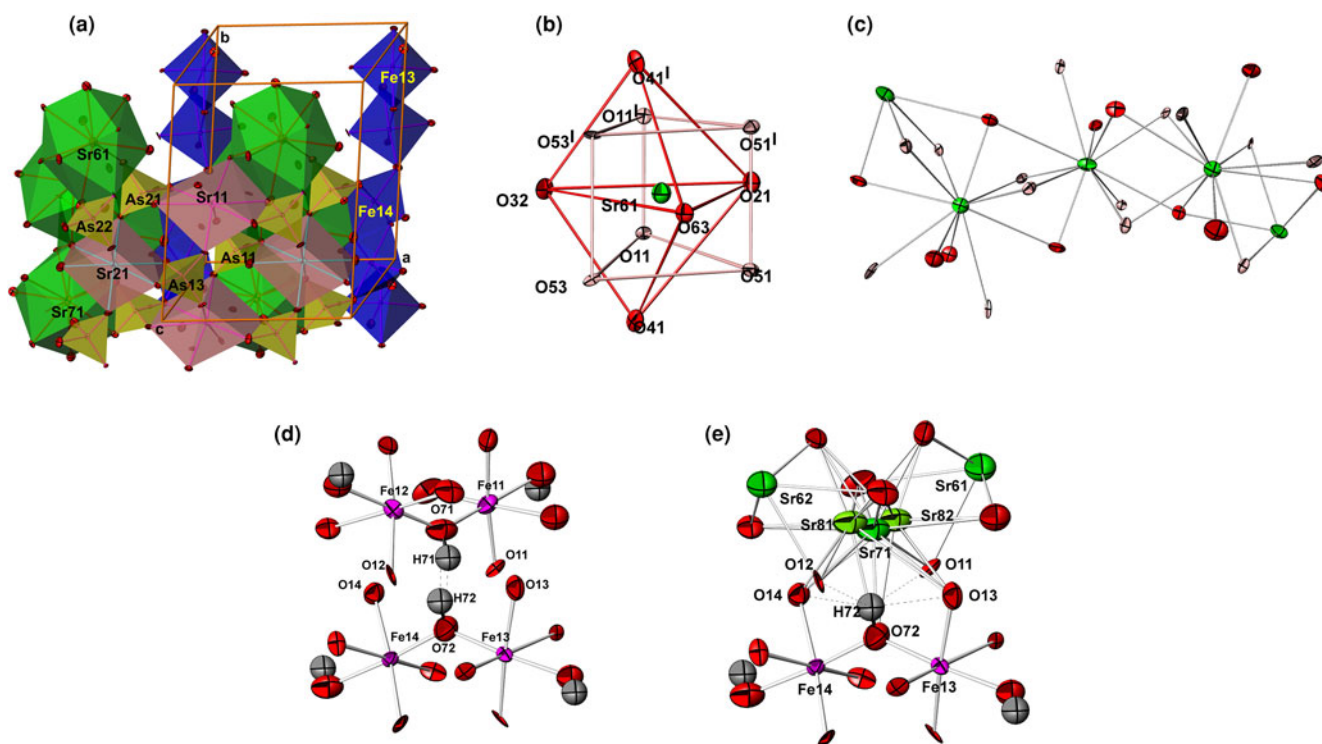


Fig. 3. (a) Polyhedral view of **II** showing two octahedral $\text{Fe}(\text{O},\text{OH})_6$ -chains and two Edshammar polyhedral SrO_{11} -chains arranged parallel to the b axis; (b) Edshammar polyhedron Sr_6O_{11} with the atom labelling scheme; (c) the infinite Edshammar polyhedral SrO_{11} -chains. Sr_{81} and Sr_{82} atoms are omitted for clarity; (d) hydrogen bonds between O_{72} and O_{71} in the carminite-like part of structure; and (e) hydrogen bonds between O_{72} and two oxygens, O_{12} and O_{13} , in the arsenbrackebuschite-like part of the structure. The environment of Sr_{71} shows the positions and atom-displacement parameters of Sr_{81} and Sr_{82} atoms (occupying ca. 10% of available sites) at a distance of only 0.724(6) and 0.697(5) Å from Sr_{71} , respectively.

part of structure four arsenic atoms As_{11} – As_{14} share only three oxygens with the $\text{Fe}(\text{O},\text{OH})_6$ octahedra, while the forth oxygen is bonded, besides As, only to 8-coordinated Sr. The other two arsenic atoms As_{21} and As_{22} are linked through all four vertices to As, Sr and to the adjacent octahedral chains. In the less pronounced arsenbrackebuschite part of structure only two oxygens from AsO_4 tetrahedra are shared with $\text{Fe}(\text{O},\text{OH})_6$ octahedra. Another two are linked to the 8 and 11-coordinated Sr^{2+} . In comparison to the other As–O bonds in tetrahedra, which vary from 1.657(9) to 1.692(8) Å, only As_{13} – O_{24} and As_{14} – O_{23} bond distances are slightly elongated (1.741(9) and 1.736(8) Å), indicating a disorder and the possible involvement in hydrogen bonding.

The average $\langle \text{As}–\text{O} \rangle$ bond distances (1.680, 1.682, 1.688, 1.692, 1.681 and 1.671 Å for $\langle \text{As}_{11}–\text{O} \rangle$, $\langle \text{As}_{14}–\text{O} \rangle$, $\langle \text{As}_{21}–\text{O} \rangle$ and $\langle \text{As}_{22}–\text{O} \rangle$, respectively) are in accordance, or slightly shorter, than the sum of the effective ionic radii of As^{+5} in tetrahedral coordination and O^{2-} ($0.335 + 1.4 = 1.735$ Å; Shannon, 1976).

Another type of chain in the structure of **II** involves SrO_{11} coordination polyhedra, which are known as Edshammar polyhedra because this polyhedron was first described by Edshammar (1969). Each 11-vertex Edshammar polyhedron SrO_{11} consists of a trigonal bipyramid SrO_5 (two O atoms at apices, usually at the longest, and three O atoms in the equatorial plane, typically at the shortest Sr–O distances) combined with the trigonal, rather

deformed, prism SrO_6 (Fig. 3b). Therefore, the strontium coordination polyhedron SrO_{11} can be described as a combination of a trigonal bipyramid and trigonal prism centred by the same Sr atom surrounded by 11 oxygen atoms at distances ranging from 2.567(11) to 3.102(8) Å in Sr61O_{11} , from 2.574(10) to 3.048(8) Å in Sr62O_{11} , from 2.650(8) to 3.091(4) Å in Sr71O_{11} , from 2.568(11) to 3.017(11) Å in Sr63O_{11} , from 2.561(12) to 3.035(11) Å in Sr64O_{11} and from 2.695(5) to 3.087(7) Å in Sr72O_{11} . Observed values of 2.865, 2.858, 2.844, 2.851, 2.864 and 2.861 Å for the average $\langle \text{Sr}-\text{O} \rangle$ bond distances are very similar and close to the calculated value from effective ionic radii of Sr^{2+} cations with CN = 11 ($1.4 + 1.4 = 2.8$ Å). The value of 1.4 Å for Sr^{2+} with CN = 11 was estimated from 1.36 and 1.44 Å given by Shannon (1976) for Sr^{2+} with CN = 10 and CN = 12, respectively. In addition, this is consistent with recent work (Gagné and Hawthorne, 2016). By sharing common trapezoidal faces, each Edshammar polyhedron is linked to two similar neighbouring polyhedra. Each Edshammar polyhedron SrO_{11} shares four oxygen atoms (two from a trigonal bipyramid and two from a trigonal prism) with the neighbouring SrO_{11} -polyhedron from one side, and additional four vertices from another side, in order to form trapezoidal face-sharing SrO_{11} -chains (Fig. 3c).

The partial occupancy of mixed sites Sr71/O71 and Sr72/O72 by Sr are accompanied by the appearance of extra atoms (Fig. 3e) at distances less than 1 Å from Sr71 (Sr81 and Sr82) and Sr72 (Sr83 and Sr84). The four additional 11-coordinated Sr81-Sr84 atoms are of very low occupancy (ca. 10% of available sites) and are surrounded by oxygen atoms in a similar way, like Sr71 and Sr72 atoms (Table S5). Some displacement ellipsoids of these oxygen atoms have a small problematic form indicating a disorder in this part of structure.

Similar to in **I** the 8-coordinated Sr^{2+} cations are situated in holes between the layers of octahedral chains linked by AsO_4 tetrahedra (Fig. 3a). For 8-coordinated Sr^{2+} cations the bond lengths Sr11-O , Sr21-O and Sr22-O vary in the intervals 2.507(7)–2.789(8), 2.512(11)–3.081(10) and 2.496(7)–3.090(9) Å, respectively. The average $\langle \text{Sr}^{2+}-\text{O} \rangle$ bond distances are similar: 2.663, 2.656 and 2.651 Å for $\langle \text{Sr11-O} \rangle$, $\langle \text{Sr21-O} \rangle$ and $\langle \text{Sr22-O} \rangle$, respectively, which is very close to the calculated value from effective ionic radii of 8-coordinated Sr^{2+} cations ($1.26 + 1.4 = 2.66$ Å; Shannon, 1976) and mean bond length of 2.656 Å given by Gagné and Hawthorne (2016). The Sr11O_8 polyhedron, having four shorter and four longer Sr–O distances, can be described as an intermediate between the considerably distorted orthorhombic prism and the tetragonal elongated double split octahedron with the four axial bonds longer ($\langle 2.795 \rangle$ Å) than the four equatorial bonds ($\langle 2.531 \rangle$ Å). Another two strontium atoms display a bicapped octahedral geometry with six shorter ($\langle 2.543 \rangle$ and $\langle 2.540 \rangle$ Å) and two longer bonds ($\langle 2.993 \rangle$ and $\langle 2.984 \rangle$ Å in Sr21O_8 and Sr22O_8 , respectively). The four next-nearest oxygen atoms are more than 3.5 Å away from the central Sr11 , Sr21 and Sr22 atoms.

The difference-Fourier map did not reveal the positions of the H atoms. This was expected considering their amount, as well as the substitutional and positional disorder of Sr. Therefore the hydrogen-bonding system is complex. However, bond-valence calculations (Brown and Altermatt, 1985; Brese and O'Keeffe, 1991) show that the Sr–O, Fe–O and As–O bond lengths are consistent with the presence of 8-coordinated and 11-coordinated Sr^{2+} , 6-coordinated Fe^{3+} , 4-coordinated As^{5+} and O^{2-} ($\Sigma v_{ij} = 1.988(1)$, $2.100(7)$ and $2.126(9)$ vu for Sr11 , Sr21 and Sr22 , $\Sigma v_{ij} = 3.072(3)$, $3.106(3)$, $3.007(0)$ and $2.994(1)$ vu for Fe11-Fe14 , $\Sigma v_{ij} = 5.043(3)$, $5.086(2)$, $4.987(2)$, $4.890(2)$, $5.050(2)$ and $5.083(1)$ vu for As11-As14 , As21 and As22). The bond valence

sums Σv_{ij} for O61-O64 , O71 and O72 without the contributions of H are close to 1, while that of O31-O34 are near 1.6 vu indicating that they are involved in hydrogen bonding (Table S8). Except O23 and O24 , the calculated bond valence sums for the other oxygen atoms were ≈ 2 indicating that they hardly participate in hydrogen-bonding interactions. The small undersaturation of O23 and O24 can be the result of a positional and substitutional disorder, i.e. the displacements of O23 and O24 oxygen atoms due to the substitutions of differently sized and charged ions (Sr72 and O72) and their amounts in the coordination sphere as well as the positional disorder of Sr72 .

The four OH groups (O61H61-O64H64) act as hydrogen-bond donors to the oxygen atoms O31-O34 ($D\cdots A$ distances are in the range 2.721(10)–2.748(8) Å and $D-H\cdots A$ angles are from 151(4)–155(5)°. Another two OH groups (O71H71 and O72H72) are alternately single hydrogen bond donors and acceptors to one another ($D\cdots A$ distance is 2.695(5) Å and $D-H\cdots A$ angles are 147(3) and 145(3)°, respectively). All these hydrogen bonds are of average strength (Table 7).

It should be kept in mind that O72 has 97% occupancy, O71 has 63% and that they can be single hydrogen-bond donors and acceptors to one another in a maximum 50% of unit cells. If O71 and O72 are always protonated, which oxygen atoms can act as hydrogen-bond acceptors? The nearest neighbours of O72 and O71 are O11-O14 (interatomic distances O72-O12 , O72-O11^{vii} , O71-O13^x and O71-O14 are 2.969(11), 2.979(11), 3.042(9) and 3.032(9) Å; symmetry codes: (vii) $x, y, z-1$; (x) $x, y, z+1$) indicating possible weak hydrogen bonds. The bond valence sums for O11-O14 are 2.102, 2.111, 2.087 and 2.052 vu. The slight oversaturation at the O11-O14 sites, coupled with the very small contributions of hydrogen bonds, could be interpreted as due to disorder and limitations of bond-valence analysis (Figs 3d, e).

Crystal structure of $\text{SrFe(AsO}_4\text{)(AsO}_3\text{OH)}$ (**III**).

In contrast to **I** and **II**, in the structure of **III** the FeO_6 octahedra are completely isolated from one another by the surrounding arsenate tetrahedra. This arrangement produces unique metal arsenate octahedral-tetrahedral-quadruple chains, which are composed of one central double and two peripheral single chains extended along the a axis (Fig. 4a). The central double chain is built up of alternating corner-sharing Fe1O_6 octahedra and As1O_4 tetrahedra. The two sides of the double chain are 'out of phase' meaning that Fe1O_6 octahedra in one side are arranged in a line with, and are linked to, As1O_4 tetrahedra on the other side. The peripheral single chains are made up of alternating corner-sharing Fe2O_6 octahedra and As2O_4 tetrahedra, which are also 'out of phase' with regard to the central double chain. These octahedral-tetrahedral-quadruple chains are centred around the inversion centre at (0,0,0) with one single and one half of a double chain on each side (Fig. 4b). The chains are further linked by As3O_4 tetrahedra to form a slab parallel to the (010) plane. A slab is placed between the relative heights, y , from approximately -0.35 to 0.35 (Fig. 4b). The three-dimensional structure is formed by interconnected slabs through bridging As4O_4 tetrahedra and hydrogen bonds. The adjacent slabs are linked additionally to each other *via* bonds to Sr^{2+} cations, which are accommodated inside the channels of the hetero-polyhedral three-dimensional open framework.

The slightly distorted Fe1O_6 and Fe2O_6 octahedra share all six corners with adjacent arsenate tetrahedra. The individual Fe1-O bond lengths vary from 1.926(9) to 2.107(10) Å, while Fe2-O bond lengths vary from 1.948(9) to 2.126(9) Å (Table 5). The

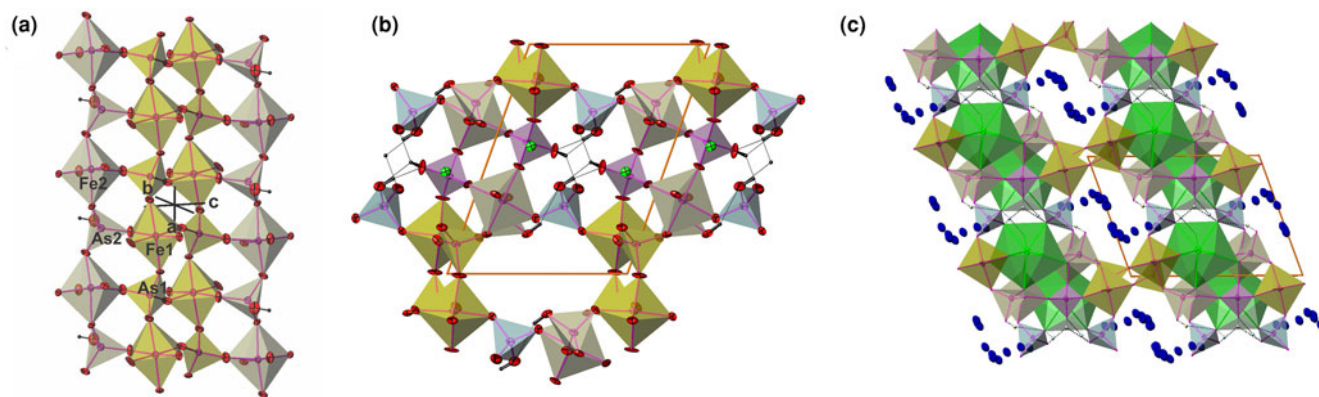


Fig. 4. Projection of a single octahedral-tetrahedral-quadruple chain showing 4M rings generated from the vertices of AsO_4 tetrahedra and FeO_6 octahedra. Double central chain is yellow, single-peripheral chains are grey. H atoms are presented as small spheres of arbitrary radii; (b) the linkages of quadruple chains into slabs projected along [100] direction (b is horizontal); (c) projection of **III** along the a -axis showing the channels where Sr2 (small blue ellipsoids) is statistically positioned in 20 sites. The Sr_2O_{16} dimers are green (c is horizontal); H atoms are presented as small spheres and H bonds are shown as broken lines.

average $\langle \text{Fe}^{3+}\text{-O} \rangle$ bond distances are similar: 2.016 and 2.002 Å for $\langle \text{Fe1-O} \rangle$ and $\langle \text{Fe2-O} \rangle$, respectively. These average values are close to the calculated from effective ionic radii of 6-coordinated Fe^{3+} ions ($0.65 + 1.4 = 2.05$ Å; Shannon, 1976).

Although the mean $\langle \text{As-O} \rangle$ bond lengths are almost the same for all arsenate tetrahedra ($\langle \text{As1-O} \rangle = 1.679$, $\langle \text{As2-O} \rangle = 1.683$, $\langle \text{As3-O} \rangle = 1.682$ and $\langle \text{As4-O} \rangle = 1.686$ Å), the individual bond lengths vary from 1.657(10) to 1.698(10), from 1.657(9) to 1.714(10), from 1.653(9) to 1.722(10) and from 1.655(9) to 1.720(10) Å for As1-O, As2-O, As3-O and As4-O, respectively. These average distances are close to the sum of the effective ionic radii of As^{5+} in tetrahedral coordination and O^{2-} ($0.335 + 1.4 = 1.735$ Å; Shannon, 1976).

Strontium cations are also linked by sharing O and OH groups. The Sr1 atoms are located inside the small channels extending nearly along [010]. The channels are centred on the inversion centre at 0.5, 0.5, 0.5 and are bounded by octahedral-tetrahedral chains with As_3O_4 and As_4O_4 tetrahedra. By sharing a common two symmetry-equivalent oxygen atom O23, two 9-coordinated Sr1 form a Sr_2O_{16} dimer. The nine Sr-O distances range from 2.473(10) to 3.120(9) Å.

The Sr2 atoms inside the large channels along [100] are highly mobile in the crystal lattice. They appear distributed among ten positions in the channels. These Sr21-Sr30 positions are too close to be occupied simultaneously and therefore Sr2 atoms are actually distributed randomly over ten pairs of symmetry-equivalent sites. The site occupancy factors of the four most occupied positions are refined to 0.210(6), 0.225(6), 0.140(6) and 0.113(6). The other positions have less than 10% occupancy.

As expected, the As-OH bond distances are slightly elongated in comparison to the other As-O bonds in tetrahedra. One OH group (O24H24) acts as a hydrogen-bond donor to the oxygen O44 from the other OH group ($D\cdots A$ distance is 3.114(16) Å and $D\cdots H\cdots A$ angle is 153°). Another two hydrogen bonds are bifurcated. The H-bond donor O34 is bound to two H-bond acceptors, O43 and O44, at $D\cdots A$ distances of 2.631(16) and 2.833(15) Å, respectively. Both are bent with $D\cdots H\cdots A$ angles of 137 and 128°. Similarly, the other H-bond donor O44 is bound to two H-bond acceptors, O34 and O43, at the comparable $D\cdots A$ distances of 2.833(15) and 3.123(14) Å, respectively, which are also bent ($D\cdots H\cdots A$ angles are 152 and 125°). Because hydrogen atoms H34 and H44 were found in general positions at a short distance of 1.48 Å their

occupancy factors were fixed at 0.5. All these hydrogen bonds are of average to weak strength (Table 8).

Bond-valence calculations (Brown and Altermatt, 1985; Brese and O'Keeffe, 1991) show that the Fe-O, As-O and Sr-O bond lengths are consistent with the presence of Fe^{3+} , As^{5+} , Sr^{2+} and O^{2-} $\Sigma v_{ij} = 3.028(1)$ vu for Fe1, $\Sigma v_{ij} = 3.151(5)$ vu for Fe2 with CN = 6, $\Sigma v_{ij} = 5.085(1)$, $\Sigma v_{ij} = 5.032(1)$, $\Sigma v_{ij} = 5.040(1)$ and $\Sigma v_{ij} = 4.986(2)$ vu for As1, As2, As3 and As4, respectively, with CN = 4 and $\Sigma v_{ij} = 2.024(2)$ vu for Sr^{2+} with CN = 10. Taking into account the contribution of the non-hydrogen atoms excluding Sr2, the oxygens, which are included in hydrogen bonding, are, as expected, undersaturated: Σv_{ij} are 1.154, 1.311, 1.473 and 1.418 vu for O24, O34, O43 and O44, respectively (Table S9). Keeping in mind that O24 is a single hydrogen bond donor towards O44, O43 is a double acceptor from O34 and O44 and O34 and O44 are alternately single hydrogen bond donors and acceptors to one another, the bond valence values are very well balanced. The bond valence sums for other oxygen atoms are very close to the nominal valence of 2⁻, or they are slightly lower indicating that they are bonded to disperse Sr2.

Raman spectroscopy

In all SrM -arsenate hydrates ($M = \text{Ni}^{2+}$ and Fe^{3+}) presented in this study, Raman bands in the 1000–600 cm^{-1} region are assigned to the symmetric and antisymmetric stretching modes of the $(\text{AsO}_4)^{3-}$ and/or $(\text{AsO}_3\text{OH})^{2-}$ groups, whereas internal bending vibrations of these tetrahedra are observed below 550 cm^{-1} , were they are partially overlaid by various external and lattice modes.

Raman spectra of **I** are shown on Fig. 5 together with Raman spectra of $\text{SrCo}_2(\text{AsO}_4)(\text{AsO}_3\text{OH})(\text{OH})(\text{H}_2\text{O})$ (Mihajlović and Effenberger, 2004) and Raman spectra of the mineral tsumcorite, $\text{PbZn}_2(\text{AsO}_4)_2 \cdot 2\text{H}_2\text{O}$ (RRUFF database: R100185, <https://rruff.info/>). $\text{SrCo}_2(\text{AsO}_4)(\text{AsO}_3\text{OH})(\text{OH})(\text{H}_2\text{O})$ was synthesised in the scope of previous work of the first author of this paper, but its Raman spectra are presented here for the first time (vs = very strong, s = strong, m = medium, w = weak, sh = shoulder). In the 1000–700 cm^{-1} range, **I** shows the As-O stretching modes of the $(\text{AsO}_4)^{3-}$ groups at: 870 (sh,s), 828 (vs), 776 (m) and 741 (w) cm^{-1} , which compare very well with the spectra of isostructural mineral tsumcorite, having its characteristic Raman bands at 870 (vs), 827 (s) and 734 (s) cm^{-1} . The Raman spectra

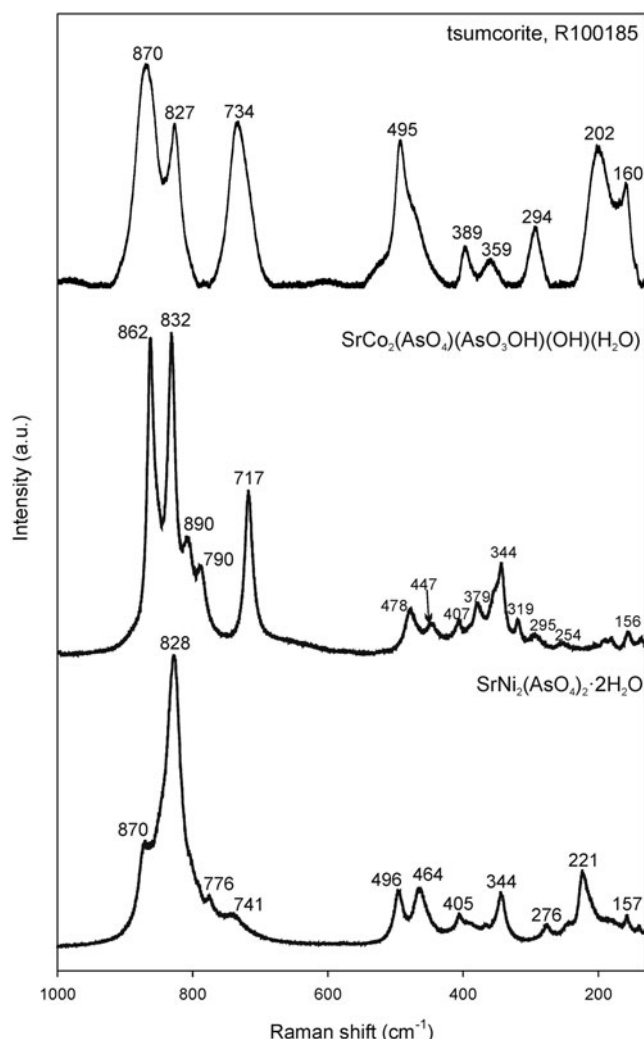


Fig. 5. Raman spectra of **I**, compared with Raman spectra of the mineral tsumcorite (RRUFF ID: R100185) and SrCo₂(AsO₄)(AsO₃OH)(OH)(H₂O) (Mihajlović and Effenberger, 2004).

of the SrCo-arsenate are more complex due to the lowering of the space-group symmetry to $P2_1/a$ and the introduction of the protonated (AsO₃OH)²⁻ group. In the As–O stretching region the SrCo-arsenate shows characteristic Raman bands at 862 (vs), 832 (vs), 890 (sh, m), 790 (sh, m) and 717 (s) cm⁻¹.

Raman spectra of **II** are characterised with the As–O symmetric and antisymmetric stretching modes of the (AsO₄)³⁻ groups at 916 (sh, w), 864 (vs), 851 (vs) and 814 (s) cm⁻¹. The band at 740 (m) cm⁻¹ may be attributed to the hydroxyl deformation mode (Fig. 6). They are compared to the Raman spectra of structurally related carminite, PbFe³⁺(AsO₄)₂(OH)₂ (RRUFF data base: R061061; <https://rruff.info/>) and arsenbrackebuschite, Pb₂(Fe³⁺, Zn)(AsO₄)₂(OH, H₂O) (RRUFF database: R100184; <https://rruff.info/>). In the region 1000–600 cm⁻¹, the new compound **II** is more similar to carminite, then brackebuschite, which was also confirmed by the structural analysis.

Raman spectra of **III** reflect the complexity of the crystal structure. The considerably large number of bands, which cannot be certainly specified, are caused by the four crystallographically different AsO₄ tetrahedra (of which three are protonated AsO₃OH groups). Spectral data on orthoarsenates that have been published

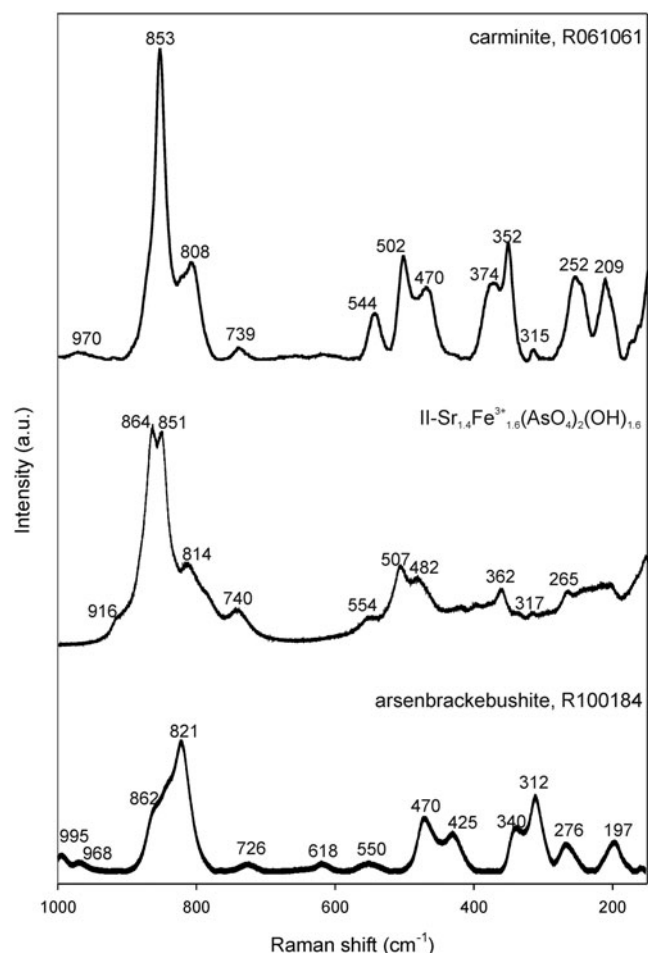


Fig. 6. Raman spectra of **II**, compared with Raman spectra of the carminite (RRUFF ID: R061061) and arsenbrackebuschite (RRUFF ID: 100184).

previously are, so far, somewhat incomplete and not in good agreement with each other. Therefore, attempts to compare with them failed. However, the distinct frequency ranges may be assigned as follows:

Based on the d - ν correlation for hydrogen bonds (Libowitzky, 1999), the Raman-shift values observed in the O–H stretching region are in very good agreement with the refined O–H...O bond lengths ranging between 2.63(2) and 3.12(1) Å (corresponding Raman bands obtained between 2500 and 3700 cm⁻¹) (Fig. 7).

The 1000–700 cm⁻¹ range shows the As–O stretching modes of the (AsO₃OH)²⁻ and (AsO₄)³⁻ groups at: 934 (sh, m), 915 (vs), 901 (sh, m), 881 (m), 40 (sh, s), 824 (s), 790 (sh, m), 775 (m) and 733 (w) cm⁻¹. The most intense Raman bands at ~915, 840 and 824 cm⁻¹ correspond to the stretching vibrations of (AsO₃OH)²⁻ and (AsO₄)³⁻ groups.

Structural relations to the associated minerals and inorganic compounds

SrNi₂(AsO₄)₂·2H₂O (**I**), is isostructural with minerals of the tsumcorite group, which adopt the monoclinic symmetry, space group $C2/m$ ($Z=2$). The general formula of the tsumcorite group is $A(M^{2+,3+})_2(XO_4)_2(H_2O, OH)_2$ (A is a large cation, $M^{2+,3+}$ is usually an octahedrally coordinated transition metal cation and X is a tetrahedrally coordinated small cation). The

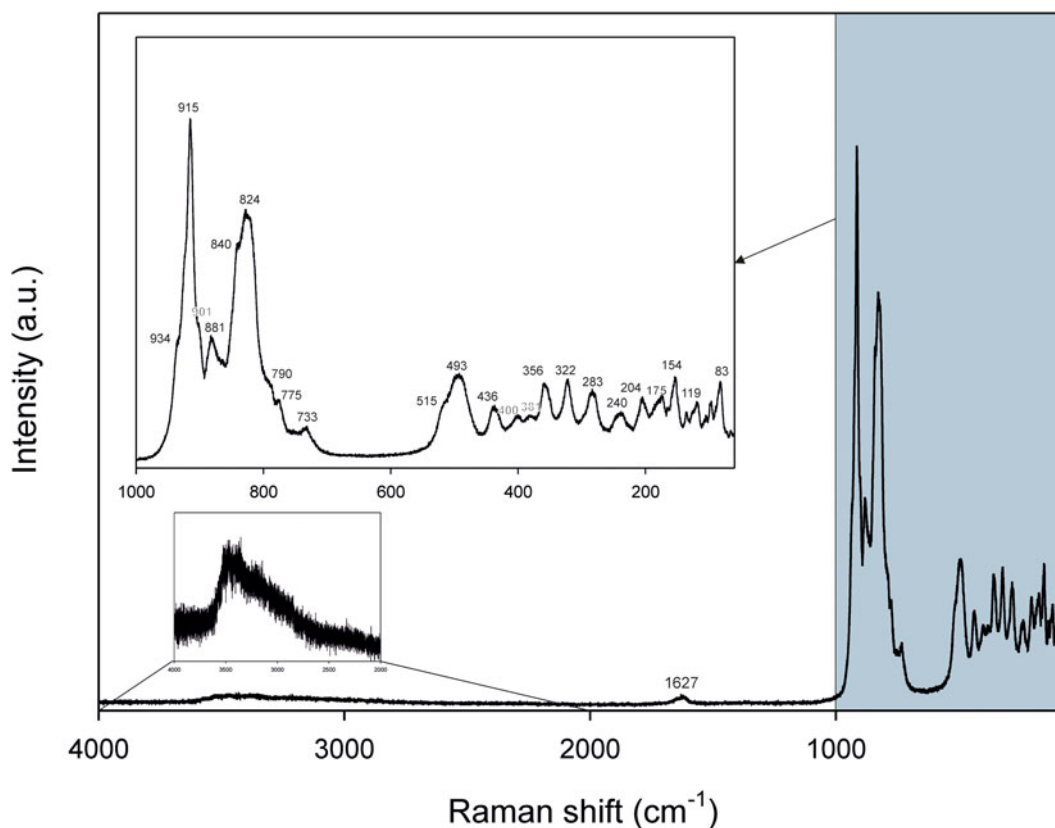


Fig. 7. Raman spectrum of III.

minerals of this group have been described with Pb, Ca, Bi and Na on the A site, Zn, Fe^{3+} , Cu^{2+} , Mn^{3+} and Co on the M site, and As^{5+} , P^{5+} , V^{5+} and S^{6+} on the X site (Tillmanns and Gebert, 1973; Pring *et al.*, 1989; Kharisun *et al.*, 1997; Krause *et al.*, 1998, 2001a, 2001b, 2002; Effenberger *et al.*, 2000; Brugger *et al.*, 2000, 2001, 2002; Yang *et al.*, 2012; Elliott and Pring, 2015; Pekov *et al.*, 2016). They display either monoclinic or triclinic symmetry, depending on the chemical composition as well as ordering of the hydrogen bonds and cations on the M sites. The basic tsumcorite-type structure mentioned previously is monoclinic (space group $C2/m$, $Z=2$) and represents the parental structure of the group referred to as tsumcorite-type structure. It was determined by Tillmanns and Gebert (1973). The majority of the tsumcorite-group minerals (Table 9) crystallise in space group $C2/m$, but some members display lower symmetry. The next modification is also monoclinic referred to as the Sr-Co-type structure. It is characterised by protonated anion tetrahedra (Mihajlović and Effenberger, 2004) and a doubled cell volume ($Z=4$). In addition, the space-group symmetry is lowered to the $P2_1/a$. The triclinic members of the tsumcorite group (Table 9) crystallise in space group $P\bar{1}$ ($Z=1$) and can be divided into the helmutwinklerite and rappoldite subgroup (Effenberger *et al.*, 2000). In the helmutwinklerite subgroup, referred to as a helmutwinklerite-type structure, the symmetry reduction is due to the ordering of the hydrogen bonds. The characteristic of this subgroup is two H_2O per formula unit and the hydrogen atoms are rearranged in order to avoid the symmetry-restricted hydrogen bonds. These changes in hydrogen bonding induce structural transformations that cause the transition from the

monoclinic tsumcorite-type structure to the triclinic helmutwinklerite-type structure (Krause *et al.*, 1998, 2001b; Effenberger *et al.*, 2000; Elliott and Pring, 2015). In the second triclinic subgroup, referred to as the gartrellite type, symmetry reduction is due to the cation ordering on the two M positions in the triclinic unit cell.

The compound II, $\text{Sr}_{1.4}\text{Fe}_{1.6}(\text{AsO}_4)_2(\text{OH})_{1.6}$ adopts a novel intermediate crystal structure between carminite and arsenbrackebuschite, nominally $\text{PbFe}_2^{3+}(\text{AsO}_4)_2(\text{OH})_2$ and $\text{Pb}_2\text{Fe}^{3+}(\text{AsO}_4)_2(\text{OH})$, respectively (Donaldson *et al.* 1955; Finney, 1963; Hofmeister and Tillmanns, 1978; Olmi and Sabelli, 1995; Kharisun *et al.*, 1996; Foley *et al.*, 1997; Gonzalez del Tanago *et al.*, 2003; Sanjeeva *et al.*, 2015; Lafuente and Downs, 2016; Table 10). Carminite is also related structurally to the phosphate analogue crimsonite, $\text{Pb}^{2+}\text{Fe}_2^{3+}(\text{PO}_4)_2(\text{OH})_2$ and the Ca analogue sewardite, $\text{Ca}^{2+}\text{Fe}_2^{3+}(\text{AsO}_4)_2(\text{OH})_2$ (Kampf *et al.*, 2016; Roberts *et al.*, 2002). Mawbyite, nominally $\text{Pb}^{2+}\text{Fe}_2^{3+}(\text{AsO}_4)_2(\text{OH})_2$, is a dimorph of carminite (the same formula, but different structure) and Fe analogue of tsumcorite (Kharisun *et al.*, 1997; Pring *et al.*, 1989, Table 10). One of the main differences between these three structures is the A: M^{3+} ratio, which is 1:2 for carminite, 1.00:1.14 for II and 1.0:0.5 for arsenbrackebuschite. These structures are also related to previously mentioned tsumcorite, nominally $\text{PbFe}_2^{2+}(\text{AsO}_4)_2 \cdot 2\text{H}_2\text{O}$, both in chemical composition and in structure properties. Carminite and tsumcorite have the same general formula $A(M^{2+,3+})_2(\text{XO}_4)_2(\text{H}_2\text{O},\text{OH})_2$. The changing of the cation valence and size in the octahedral sites from basically M^{3+} in carminite to M^{2+} in tsumcorite causes changes in the octahedral chains, which are the main building units. The presence of

Table 9. Comparison of the unit-cell parameters of the compounds adopting the tsumcorite-type of structure, $A(M^{2+}, M^{3+})_2(XO_4)_2(H_2O, OH)_2$ (s.g. $C2/m$, $Z = 2$ for 1–18, s.g. $P2_1/a$, $Z = 4$ for 19 and s.g. $P\bar{1}$, $Z = 1$ for 20–25).

	Formula, mineral name	<i>a</i> (Å)	<i>b</i> (Å)	<i>c</i> (Å)	$\alpha / \beta / \gamma$ (°)	Volume (Å ³)	Reference
1	Pb(Zn _{0.5} Fe _{0.12} Fe _{0.38} ³⁺)(AsO ₄) ₂ [(OH) _{0.38} (H ₂ O) _{0.62}] ₂ <i>tsumcorite</i>	9.124(3)	6.329(2)	7.577(2)	115.28(3)	395.5(4)	Tillmanns and Gebert (1973)
2	Pb(Co _{0.39} Fe _{0.37} Ni _{0.20} Zn _{0.02} Al _{0.02})(AsO ₄) ₂ [(OH) _{0.39} (H ₂ O) _{0.61}] ₂ <i>cobalttsumcorite</i>	9.097(2)	6.313(2)	7.555(3)	115.08(2)	393.0(1)	Krause et al. (2001a)
3	(Pb _{0.7} Ca _{0.3})(Ni _{1.36} Fe _{0.49} Zn _{0.06} Mn _{0.04} Mg _{0.04} Co _{0.01})(AsO ₄) ₂ [(OH) _{0.53} (H ₂ O) _{0.47}] ₂ <i>nickeltsumcorite</i>	9.124(8)	6.339(3)	7.567(7)	115.19(6)	396.0(7)	Pekov (2016)
4	Pb(Fe _{1.5} Zn _{0.5})(AsO ₄) ₂ [(OH) _{1.5} (H ₂ O) _{0.5}] ₂ <i>mawbyite</i>	9.052(5)	6.277(4)	7.580(3)	114.57(3)	391.6(4)	Pring et al. (1989)
5	Pb(Fe _{1.93} Zn _{0.07})(AsO ₄) ₂ [(OH) _{1.93} (H ₂ O) _{0.07}] ₂ <i>mawbyite</i>	9.066(4)	6.286(3)	7.564(3)	114.857(5)	391.1(3)	Kharisun et al. (1997)
6	Ca(Zn _{1.02} Mn _{0.98})(AsO ₄) ₂ [(OH) _{0.49} (H ₂ O) _{0.51}] ₂ <i>lotharmeyerite</i>	9.0727(6)	6.2530(4)	7.4150(5)	116.739(4)	375.68(4)	Yang et al. (2012)
7	Ca(Fe _{0.99} Zn _{1.01})(AsO ₄) ₂ [(OH) _{0.50} (H ₂ O) _{0.50}] ₂ <i>ferrilotharmeyerite</i>	9.010(3)	6.246(2)	7.391(2)	115.52(3)	375.36	Krause et al. (1998)
8	Ca(Mn _{0.59} □ _{0.24} Mg _{0.17})(AsO ₄) ₂ [(OH) _{0.76} (AsO ₂ (OH) ₂] ₂ [(OH) _{0.59} (H ₂ O) _{0.41}] ₂ <i>manganlotharmeyerite</i>	9.043(1)	6.2314(7)	7.3889(9)	116.392(2)	372.97(9)	Brugger et al. (2002)
9	(Ca _{0.97} Bi _{0.03})(Ni _{1.10} Fe _{0.70} Co _{0.20})(AsO ₄) ₂ [(OH) _{0.36} (H ₂ O) _{0.64}] ₂ <i>nickellotharmeyerite</i>	9.005(1)	6.205(1)	7.411(1)	115/31(1)	374.4(1)	Krause et al. (2001a)
10	Ca(Mg _{0.78} Al _{0.82} Fe _{0.35} Mn _{0.05})(AsO ₄) ₂ [(OH) _{0.61} (H ₂ O) _{0.39}] ₂ <i>cabalzarite</i>	8.925(2)	6.143(1)	7.352(1)	115.25(3)	364.6(1)	Brugger et al. (2000)
11	Ca(Mg _{0.82} Al _{0.84} Fe _{0.08} Mn _{0.26})(AsO ₄) ₂ [(OH) _{0.59} (H ₂ O) _{0.41}] ₂ <i>cabalzarite</i>	8.9284(5)	6.1314(5)	7.3439(6)	115.494(7)	362.9(1)	Brugger et al. (2000)
12	(Bi _{0.65} Ca _{0.35})(Co _{1.04} Ni _{0.60} Fe _{0.36})(AsO ₄) ₂ [(OH) _{0.51} (H ₂ O) _{0.49}] ₂ <i>schneebergite</i>	9.005(1)	6.211(1)	7.440(1)	115.19(1)	376.5(1)	Krause et al. (2002)
13	(Bi _{0.55} Ca _{0.45})(Ni _{1.00} Co _{0.62} Fe _{0.40})(AsO ₄) ₂ [(OH) _{0.50} (H ₂ O) _{0.50}] ₂ (H ₂ O,OH) ₂ <i>nickelschneebergite</i>	8.995(1)	6.207(1)	7.462(1)	115.00(1)	377.6(1)	Krause et al. (2002)
14	PbCu ₂ ²⁺ (As _{1.23} S _{0.77} O ₄) ₂ [(OH) _{0.38} (H ₂ O) _{0.62}] ₂ (H ₂ O,OH) ₂ <i>thometzekite</i>	9.077(8)	6.300(8)	7.661(8)	116.99(8)	390.38	Krause et al. (1998)
15	PbFe ₂ (VO ₄) ₂ F ₂ <i>mounanaite</i>	9.294(8)	6.166(7)	7.713(8)	115.57(6)	398.72	Krause et al. (1998)
16	(Pb _{0.83} Sr _{0.17})Mn ₂ (VO ₄) ₂ (OH) ₂ <i>krettnichite</i> (from powder data)	9.275(7)	6.284(3)	7.682(2)	117.97(4)	395.4(3)	Brugger et al. (2001)
17	SrNi ₂ (PO ₄) ₂ ·2H ₂ O <i>synthetic compound</i>	8.8877(3)	6.0457(3)	7.3776(3)	114.173(2)	361.66(3)	Assani et al. (2010)
18	SrNi ₂ (AsO ₄) ₂ ·2H ₂ O <i>synthetic compound</i>	9.173(3)	6.2025(12)	7.562(3)	114.41(2)	391.8(2)	This work (I)
19	SrCo ₂ (AsO ₄)(AsO ₃ OH)(OH)(H ₂ O) <i>synthetic compound</i>	9.139(2)	12.829(3)	7.522(2)	114.33(3)	803.6(3)	Mihajlović and Effenberger (2004)
20	PbZn ₂ (AsO ₄) ₂ (H ₂ O) ₂ <i>helmutwinklerite</i>	5.606(2)	5.610(2)	7.617(1)	70.19(2) / 69.91(2) / 69.18(2)	203.77	Krause et al. (1998)
21	Pb(CoNi _{0.64} Zn _{0.36})(AsO ₄) ₂ (H ₂ O) ₂ <i>rappoldite</i>	5.595(1)	5.572(1)	7.593(1)	70.19(1) / 70.41(1) / 69.23(1)	201.91	Effenberger et al. (2000)
22	Pb(CuFe ³⁺)(AsO ₄) ₂ [(OH) _{0.50} (H ₂ O) _{0.50}] ₂ <i>gartrellite</i>	5.454(1)	5.653(1)	7.579(1)	67.52(1) / 69.27(1) / 70.08(1)	196.15	Krause et al. (1998)
23	Pb(Zn _{1.20} Cu _{0.40} Fe _{0.40} ³⁺)(AsO ₄) ₂ [(OH) _{0.40} (H ₂ O) _{0.60}] ₂ <i>zincgartrellite</i>	5.550(1)	5.620(1)	7.621(1)	68.59(1) / 69.17(1) / 69.51(1)	200.10	Effenberger et al. (2000)
24	PbCuAl(AsO ₄) ₂ (OH, H ₂ O) ₂ <i>yanowinnaite</i>	5.444(2)	5.640(2)	7.518(2)	67.89(2) / 69.48(2) / 70.18(2)	194.51(7)	Elliott and Pring (2015)
25	CaCuFe ³⁺ (AsO ₄) ₂ [(OH) _{0.50} (H ₂ O) _{0.50}] ₂ <i>lukrahnite</i>	5.457(3)	5.539(4)	7.399(6)	68.43(5) / 68.90(4) / 69.44(5)	187.8(1)	Krause et al. (2001b)

Table 10. Comparison of the unit-cell parameters of the compounds adopting carminite $[A(M^{2+}, M^{3+})_2(XO_4)_2(H_2O, OH)_2]$, (s.g. *Cccm*, $Z=8$ for 1–7) and arsenbrackebuschite $[A_2(M^{2+}, M^{3+})(XO_4)_2(H_2O, OH)]$, (s.g. $P2_1/m$, $Z=2$ for 10–15) structure type.

	Formula, mineral name	<i>a</i> (Å)	<i>b</i> (Å)	<i>c</i> (Å)	$\alpha / \beta / \gamma$ (°)	Volume (Å ³)	Reference
1	PbFe ₂ ³⁺ (AsO ₄) ₂ (OH) ₂ <i>carminite</i>	16.595(5)	7.580(5)	12.295(5)	90	1546.59	Finney (1963)
2	PbFe ₂ ³⁺ (AsO ₄) ₂ (OH) ₂ <i>carminite</i>	16.604(2)	7.586(1)	12.274(2)	90	1546.01	Olmi and Sabelli (1995)
3	PbFe ₂ ³⁺ (AsO ₄) ₂ (OH) ₂ <i>carminite</i>	16.617(4)	7.585(2)	12.272(2)	90	1546.76	Olmi and Sabelli (1995)
4	PbFe ₂ ³⁺ (AsO ₄) ₂ (OH) ₂ <i>carminite</i>	16.602(3)	7.586(1)	12.268(3)	90	1545.07	Olmi and Sabelli (1995)
5	PbFe ₂ ³⁺ (AsO ₄) ₂ (OH) ₂ <i>carminite</i>	16.591(2)	7.580(1)	12.285(1)	90	1544.96	Kharisun <i>et al.</i> (1996)
6	CaFe ₂ ³⁺ (AsO ₄) ₂ (OH) ₂ <i>sewardite</i>	16.461(2)	7.434(1)	12.131(2)	90	1484.48	Roberts <i>et al.</i> (2002)
7	(Pb _{1.06} Ca _{0.06})(Fe _{1.71} Zn _{0.05} Al _{0.04} Cu _{0.01}) (P _{1.32} As _{0.61} V _{0.05} S _{0.02})O ₈ [(OH) _{1.64} (H ₂ O) _{0.36}] <i>crimsonite</i>	16.2535(13)	7.4724(4)	12.1533(9)	90	1476.04(17)	Kampf <i>et al.</i> (2016)
8	Sr _{1.4} Fe _{1.6} (AsO ₄) ₂ (OH) _{1.6} <i>synthetic compound</i>	9.072(2)	12.227(2)	7.546(2)	114.575(6)	761.2(2)	This work (II)
9	Pb ₂ (Fe _{0.65} Zn _{0.35})(AsO ₄) ₂ [(OH) _{0.65} (H ₂ O) _{0.35}] <i>arsenbrackebuschite</i>	7.763(1)	6.046(1)	9.022(1)	112.5(1)	391.2(1)	Hofmeister and Tillmanns (1978)
10	Ba ₂ Fe ³⁺ (AsO ₄) ₂ (OH) ₂ <i>canosioite</i>	7.8642(4)	6.1083(3)	9.1670(5)	112.874(6)	405.73(4)	Cámara <i>et al.</i> (2017)
11	Pb ₂ (Mn, Fe)(VO ₄) ₂ (OH) <i>brackebuschite</i> , B ₂₁ /m	7.68(1)	6.15(5)	16.52(2)	93.75	779.525	Donaldson <i>et al.</i> (1955)
12	Pb ₂ Mn ³⁺ (VO ₄) ₂ (OH) <i>brackebuschite</i>	7.650(1)	6.135(1)	8.863(1)	111.76(1)	386.325	Foley <i>et al.</i> (1997)
13	Pb ₂ Mn ³⁺ (VO ₄) ₂ (OH) <i>brackebuschite</i>	7.6492(1)	6.1262(1)	8.9241(2)	112.195(1)	387.202(12)	Lafuente and Downs (2016)
14	Pb ₂ Fe(VO ₄) ₂ (OH) <i>calderónite</i>	7.649(1)	6.101(1)	8.904(1)	112.23(1)	384.6	Gonzalez del Tanago <i>et al.</i> (2003)
15	Sr ₂ V(VO ₄) ₂ (OH) <i>synthetic compound</i>	7.6249(15)	6.0903(12)	8.8153(18)	112.15(3)	379.15(15)	Sanjeeva <i>et al.</i> (2015)
16	SrFe(AsO ₄)(AsO ₃ OH) <i>synthetic compound</i>	6.6379(13)	9.6849(19)	12.666(3)	108.86(3) / 90.97(3) / 104.49(3)	741.8(3)	This work (III)

OH or H₂O depends on the amount of M³⁺ and M²⁺ cations and the valence and quantity of the large cation A.

The novel-type crystal structure of **III** is characterised by the unique octahedral–tetrahedral–quadruple, [Fe(AsO₄)]_n infinite chains, which are composed of one central double and two peripheral single chains extended along the *a* axis. Similar vertices-sharing double-sided chains of octahedra and tetrahedra has been found in galliskiite, Ca₄Al₂(PO₄)₂F₈·5H₂O (Kampf *et al.*, 2010) and some phosphates (Mills *et al.*, 2010a, 2010b and references therein). Dissimilar quadruple chains are found in the synthetic ferric sulfate trihydrate, Fe₂(SO₄)₃·3H₂O (Xu and Parise, 2011). Besides the exceptional occurrence of quadruple chains of FeO₆ and AsO₄ polyhedra, the characteristics of this structure are the novel chemical composition and the extremely high positional disorder of Sr₂ atoms.

Conclusion

The results of phase formation studies in the system Sr–M²⁺–As–O–H increase the knowledge on SrM-oxoarsenates(V) and their crystal structures. All three structures have different chains as the primary building units. Compound **I** is a new strontium nickel analogue of tsumcorite PbZn₂(AsO₄)₂·2H₂O (monoclinic, s.g. *C2/m*) having a structure which is built up of infinite linear edge-sharing NiO₄(OH)₂ octahedral chains with AsO₄ tetrahedra and SrO₈ polyhedra linking them by sharing corners. Compound **II** exists as an unusual intermediate form between carminite, PbFe₂³⁺(AsO₄)₂(OH)₂ and arsenbrackebuschite, Pb₂Fe³⁺(AsO₄)₂(OH) type structures. The characteristic feature of compound **II** is that it hosts two type of chains: (i) octahedral vertex- and edge-sharing FeO₄(OH)₂ chains like in carminite and (ii) chains similar to arsenbrackebuschite which involves SrO₁₁ coordination polyhedra, known as Edshammar polyhedra, that can be described as a combination of a trigonal bipyramid and trigonal prism centred by the same Sr atom surrounded by 11 oxygen atoms.

The MΦ₆ octahedra (Φ = O, OH or OH₂) in **I** like in tsumcorite and arsenbrackebuschite is all edge sharing, in contrast to the

alternating edge and vertices sharing in carminite and **II**. In carminite and **II** structures two MΦ₆ octahedra share edges to form a M₂Φ₁₀ dimer and the M₂Φ₁₀ dimers share corners to form an M₂Φ₈ chain that is decorated by XO₄ tetrahedra. These parallel chains are linked to form a framework by sharing octahedron and tetrahedron vertices. Contrary to this, the arsenbrackebuschite and tsumcorite MΦ₆ octahedra share one set of *trans* edges with adjacent octahedra to form a MΦ₄ chain. Adjacent octahedra are bridged by XO₄ tetrahedra to form a decorated chain. The parallel chains in tsumcorite are further linked in the M₂(XO₄)₂Φ₂ sheets similar to those in natrochalcite (nominally NaCu₂(SO₄)₂(OH)·2H₂O). No such linkage occurs in arsenbrackebuschite where the MΦ₆ octahedra occupy only half of the octahedral sites, because every second octahedral chain is omitted and replaced by chains of 11-coordinated Pb. Similar substitutions were found in **II** i.e. nearly 20% of the octahedra chains are replaced by SrO₁₁ polyhedra chains. The 8- and 11-coordinated Sr atoms in **II** adopt an almost identical polyhedral environment, such as in the synthetic compound Sr₂V(VO₄)₂(OH) with a brackebuschite-type structure (Sanjeeva *et al.*, 2015).

Compound **III** exhibits a new type of structure having interesting novel quadruple chains which are composed of FeO₆ octahedra alternating with AsO₄ tetrahedra sharing corners. In the structure of **III** the quadruple chains are connected with AsO₄ tetrahedra in a 3D open framework with the channels occupied by Sr atoms.

The chains in **I** have the same topology as the chains in tsumcorite and the chains in **II** have the topology as the chains in carminite and brackebuschite. Quadruple chains of a topology different to that of **III** were found in the synthetic compound Fe₂(SO₄)₃·3H₂O (Xu and Parise, 2011).

The further interesting feature of these structures is variable coordination geometry of Sr²⁺ cations combined with the substitutional (in **II**) and positional (in **II** and **III**) disorder. The strontium coordination geometries include near-prismatic SrO₈ coordination polyhedra in **I** and **II**, the chains of Edshammar SrO₁₁ coordination polyhedra in **II** and mutually isolated Sr₂O₁₆ dimers in **III**.

We have shown that the compounds from the $M1$ – $M2$ –As–O–H system ($M1 = \text{Ca}^{2+}$, Sr^{2+} and Ba^{2+} ; $M2 = \text{Mn}^{2+,3+}$, $\text{Fe}^{2+,3+}$, Co^{2+} , Ni^{2+} , Cu^{2+} and Zn) are not fully structurally and chemically defined. However, they could play an important role as the arsenic scavengers in arsenic-extreme environments, such as sulfidic mining wastes and tailings. Two of three investigated phases crystallises in a novel structure type and thus add details to the crystal chemistry of oxoarsenate(V) compounds.

Supplementary material. To view supplementary material for this article, please visit <https://doi.org/10.1180/mgm.2021.41>

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