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THE QUATERNARY ISOTOPE LABORATORY THERMAL DIFFUSION ENRICHMENT SYSTEM: DESCRIPTION AND PERFORMANCE

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ABSTRACT. The thermal diffusion enrichment system of the Quaternary Isotope Laboratory consists of 23 hot wire columns of 3m effective length combined to 2 separate systems of 3 and 3 separate systems of 4 columns at the top, each system in series with 1 bottom column. From ≈ 130 L NTP of CO (~ 65 g of carbon) it produces ≈ 8 L NTP of CO (~ 4 g of carbon) enriched in $^{12}\text{C}^{18}\text{O}$ by a factor 6 to 7 and in $^{14}\text{C}^{16}\text{O}$ by a factor 7 to 8 in about 5 weeks. For $^{12}\text{C}^{18}\text{O}$ the system has a theoretical equilibrium separation factor of about 250 and a theoretical equilibrium enrichment of about 15. For $^{14}\text{C}^{16}\text{O}$ these values are 1300 and 16, respectively. The dependence of thermal diffusion transport on gas exchange between top and bottom section and between columns and reservoirs and on wire temperature is given. Forced gas exchange and a higher wire temperature gave a more rapidly increasing enrichment without substantially increasing its final value of 6 to 7 for $^{12}\text{C}^{18}\text{O}$. A comparison with the Groningen enrichment system shows that the two systems behave very similarly and that not the system geometry but individual column parameters and the ratio total sample mass/enriched sample mass are the dominant factors determining the enrichment.

INTRODUCTION

In the Quaternary Isotope Laboratory at the University of Washington the radiocarbon dating range for a 4-day counting period has been increased to 60,000 years by the use of carefully constructed and shielded large counters placed in an underground measuring chamber (Stuiver, Robinson, and Yang, 1976). Another approach to an extended age range has been the construction of an enrichment system to increase by a known factor the ^{14}C activity of very old samples. This has pushed the dating range back to 75,000 years (Stuiver, Heusser, and Yang, 1978). In the apparatus, a large amount of sample in the form of CO gas (≈ 130 L NTP ~ 65 g of C) is subjected to thermal diffusion (TD) isotope separation. A smaller quantity of gas (≈ 8 L NTP ~ 4 g of C) enriched in the heavy isotopes is obtained.

The use of CO for the enrichment of ^{14}C and the relation between the simultaneous enrichments of the isobaric molecules $^{14}\text{C}^{16}\text{O}$ and $^{12}\text{C}^{18}\text{O}$, important to determine the ^{14}C enrichment, were first studied by de Vries (1956). The study resulted in the construction, in Amsterdam, of a thermal diffusion enrichment set-up used to enrich ^{14}C samples for the Groningen Radiocarbon Dating Laboratory. This extended the range of this laboratory to 70,000 years (Haring, de Vries, and de Vries, 1958; Vogel and Zagwijn, 1967; Vogel and Waterbolk, 1972). Contamination resulted in a temporary halt of the enrichment activities until, in 1970, the Amsterdam enrichment system was relocated to Groningen. After the earlier

contamination problems had been solved, a series of dates up to 73,000 years BP were obtained (Grootes *et al*, 1975; Grootes, 1977; 1978).

The use of thermal diffusion enrichment to extend the range of radiocarbon dating has also been studied in other laboratories. Concentric tube columns for the enrichment of ^{14}C in methane were constructed and studied by Dickel and coworkers in Munich (Malota, 1962; Kretner, 1973; Kretner and Dickel, 1975), by Erlenkeuser (1971a, b, 1976) in Kiel, and by Felber and Pak in Vienna (Pak, 1970; Felber and Pak, 1972). None of these systems appears to be operational for radiocarbon dating yet. The present day radiocarbon dating range is, therefore, the 75,000 years obtained at the Quaternary Isotope Laboratory and in Groningen.

The newly developed technique of direct detection of ^{14}C using an accelerator as a very sensitive mass spectrometer (Muller, 1977; Muller, Stephenson and Mast, 1978; Bennet *et al*, 1977; 1978; Nelson, Korteling, and Scott, 1977) may give a similar range when the problems of contamination and machine and source instability have been vanquished. The main advantage of this new method is that it enables age determinations on very small samples (*cf* Stuiver, 1978).

The construction of the present University of Washington (UW) system began at Yale in the late sixties. CO was chosen as the enrichment medium because (i) it is light (*ie*, the relative mass difference between $^{14}\text{C}^{16}\text{O}$ and $^{12}\text{C}^{16}\text{O}$ is large); (ii) its intermolecular interaction has a short range, which is favorable for TD separation; (iii) it has a fairly good thermal stability; (iv) it is very easy to make by passing CO_2 , the first step in the preparation of any ^{14}C sample, over hot zinc; (v) it offers the possibility to determine the ^{14}C enrichment by measuring the simultaneous enrichment of molecules of mass 30 (mainly $^{12}\text{C}^{18}\text{O}$) mass spectrometrically. The alternative, CH_4 , was less attractive because the advantage of its smaller mass is offset by a longer interaction range and less thermal stability, while its production, although also a standard procedure in many laboratories (Agrawal, Gupta, and Kusumgar, 1971; Buddemeier *et al*, 1970), is somewhat more complicated for the large samples used. CH_4 also does not readily offer a check on the enrichment. The fact that any tritium contamination incorporated in the methane will undergo an enrichment similar to that of ^{14}C is an additional complication for the dating of samples in the 50,000 to 75,000 year range.

In the UW system, the same basic columns were used as in the original de Vries system in Amsterdam, but the geometric configuration is different. In this paper, we describe the enrichment set-up and compare its performance with that of the Amsterdam-Groningen thermal diffusion system and with theory.

APPARATUS

Before an enrichment sample can be counted, it must undergo the following conversions: (i) combustion to CO_2 ; (ii) reduction to CO; (iii) enrichment; (iv) oxidation to CO_2 . We will first discuss these conversions.

Combustion of the pretreated sample takes place in a system that is only used for old samples. It is similar to the system for routine combus-

tion of ^{14}C samples (de Vries and Barendsen, 1953), but its dimensions have been changed to produce large quantities of CO_2 . Typically, the UW system produces about 100 L NTP of CO_2 in 5 hours.

In the *reduction system*, the CO_2 sample is reduced to CO by passing through a zinc-asbestos mixture (10 to 1 by weight) at about 400°C . Basic aspects of system and procedure are similar to that described by Grootes *et al* (1975). A large batch of zinc-asbestos is used which can reduce between 5 and 10 samples before replacement is needed. Typically, about 10 L NTP of CO is produced in 1 hour at a pressure difference over the zinc-asbestos of about 350 torr. The CO is condensed with liquid nitrogen (vapor pressure ≈ 470 torr at liquid nitrogen temperature) to keep the pressure in the system low. When the reduction has been completed, the CO is fed into the thermal diffusion system through a cold trap to catch any CO_2 that is still present.

The *enrichment system* consists of 23 single thermal diffusion columns; 18 columns in parallel in a top section are combined in 2 sets of 3 and 3 sets of 4 columns that are each in series with a single column in the bottom section (fig 1). Each column is of the Clusius-Dickel type and consists of a pyrex glass tube with an inner diameter of 1.0cm (top) or 0.9cm (bottom section). A 0.3mm diameter Platinum-Iridium (80%-20%) wire is suspended in the center. The effective length of the wire and column is 3.0m. The wall of the pyrex tube is cooled on the outside by water flowing up between the wall and a second concentric pyrex tube of 3cm diameter. The column construction is similar to that published by Grootes *et al* (1975), except for the lack of spacers, which facilitated construction.

The top columns are interconnected at the bottom; they must be separate at the top to prevent parasitic convection caused by small temperature differences between the columns. Therefore each top column has its own 5 L storage reservoir. The gas from these reservoirs is circulated over the top of the columns by convection. The connection between top and bottom columns forms a parasitic volume of $\approx 74\text{cm}^3$ (3 column system) and $\approx 93\text{cm}^3$ (4). Each of the bottom columns is connected to an enrichment reservoir of 1.2 L (the sets of 3) or 1.6 L (the sets of 4 columns). The gas in these reservoirs is combined with that of the bottom columns for the enriched sample. The enrichment is made at a pressure of about 850 torr. Power consumption is about 1 kW (67 V, 15A) and 0.8 kW (67 V, 12A) for the 4 and the 3 column systems, respectively, (~ 0.67 W/cm wire), resulting in a wire temperature of about 960 K. The storage and enrichment reservoir and the parasitic volume are at ambient temperature which, near the system, is ≈ 300 K. The cold wall temperature is estimated to be in the range of $25\text{-}45^\circ\text{C}$, typically 308 K. It varies with the flow rate of the cooling water (≈ 6 L/min, *ie*, 0.26 L/min/column) and with the seasonal fluctuations of the temperature of the water entering the columns (range $10\text{-}18^\circ\text{C}$). Gas properties used in TD calculations are taken at the cold wall temperature of 308 K. The average temperature of the gas in the columns is 416 K.

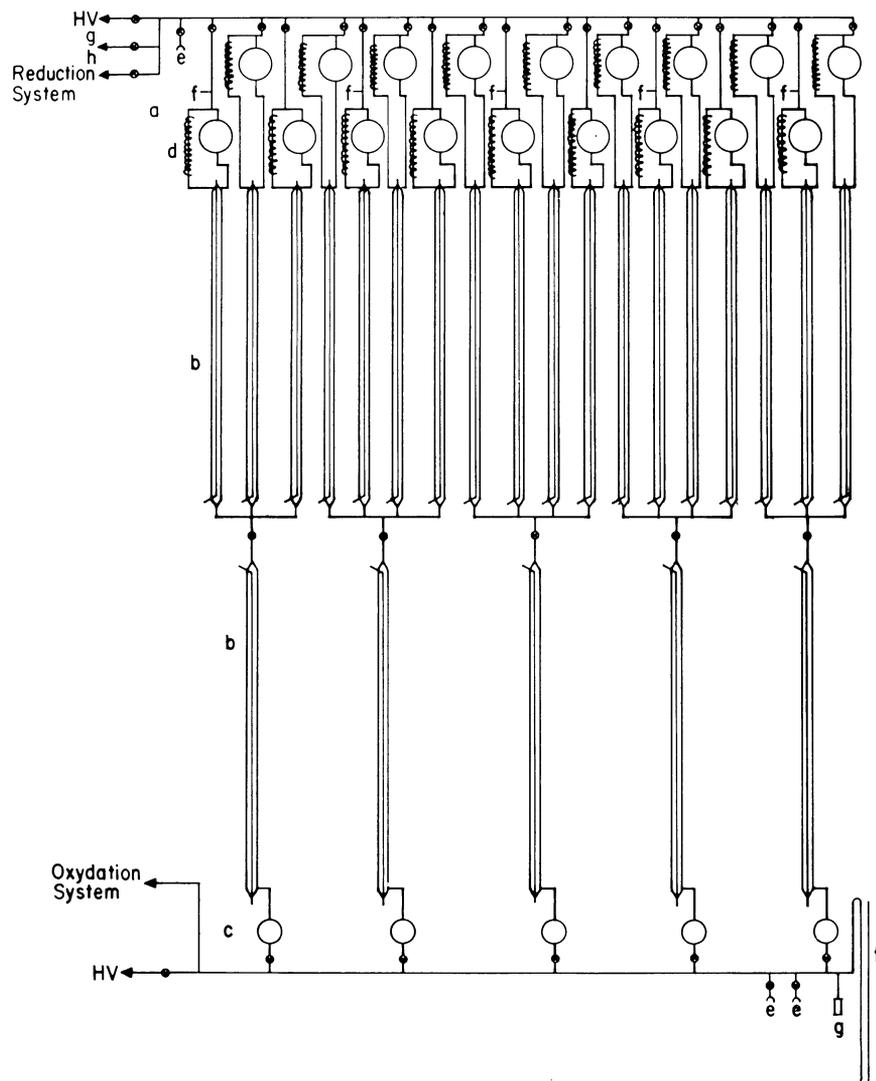


Fig. 1. Thermal diffusion enrichment system. a—storage volume V_- , b—thermal diffusion columns, c—enriched sample volume V_+ , d—thermal circulation, e—sampling point, f—mercury manometer, g—thermocouple gauge, h—Penning gauge, HV—high vacuum pump.

In the oxidation system (fig 2) the enriched CO from the enrichment reservoirs plus bottom columns is converted back to CO₂. The CO is first adsorbed on silica gel at liquid nitrogen temperature and then expanded through a CuO furnace at 450°C into a 10 L storage bulb. The gas mixture then circulates through the CuO by convection till the oxidation is complete. After about a week, the CO₂ is condensed with liquid N₂, expanding residual CO from the silica into the system, and the silica is closed off from the system. The oxidation is then resumed until completion, usually after 2 to 3 weeks. Virtually no non-condensable gas remains. Oxidation must be complete to ensure that any isotope fractionation taking place during the process does not influence the carbon isotopic composition of the sample. The oxygen isotopic composition will change because of exchange with the CuO. The enriched CO₂ sample is either cleaned further by circulation through a Cu furnace or transferred to a storage flask to wait for the radon decay (from the silica) and measurement. Before each oxidation, the CuO furnace is regenerated by heating at ≈ 450°C in oxygen. Then furnace and silica are thoroughly degassed.

Measurement: A Nuclide 6-60 RMS mass spectrometer operated in the single collector mode was used to measure the stable isotopic com-

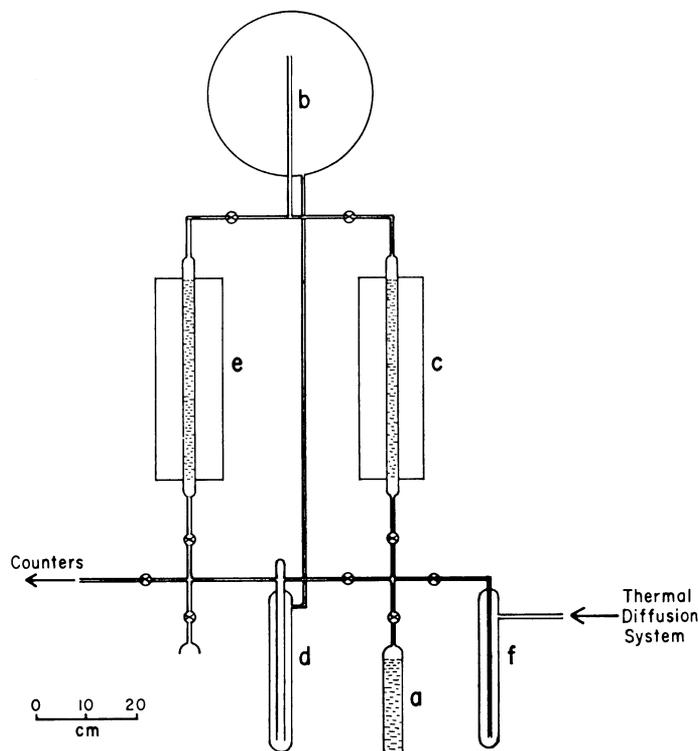


Fig. 2. Oxidation system. a—silica adsorber, b—10 L storage bulb, c—CuO furnace at 450°C, d—CO₂ trap, e—Cu furnace at 450°C, f—trap for CO₂ in enriched CO sample.

position (mass 28: $^{12}\text{C}^{16}\text{O}$; mass 29: mainly $^{13}\text{C}^{16}\text{O}$, some $^{12}\text{C}^{17}\text{O}$; mass 30: mainly $^{12}\text{C}^{18}\text{O}$, some $^{13}\text{C}^{17}\text{O}$, $^{14}\text{C}^{16}\text{O}$) of the sample. The small amount of gas needed for a mass spectrometric measurement makes it possible to sample the gas during the enrichment without seriously disturbing the system.

The ^{14}C activities were measured in a 2.3 L proportional CO_2 counter (Stuiver, Robinson, and Yang, 1976).

DISCUSSION

The first question to be asked for a new enrichment system is, what and how fast can enrichment be obtained? Secondly, how does the system performance compare with other systems and with theory? The theory is important because it may indicate what improvements are possible in final enrichment or in the rate of enrichment. Theoretical detail is avoided in this paper, the specific calculations of which the results are used here will be published elsewhere. Discussed in the following sections are (a) the working and design of a single column and that of the complete system,

TABLE I
List of symbols

C_1, C_2	= mole fraction of the abundant, light (1) and the rare, heavy (2) isotopic molecule, respectively
H	= thermal diffusion transport coefficient
H_s	= thermal diffusion transport coefficient for a system consisting of several columns in parallel and/or in series
K	= $K_c + K_d + K_p$
K_c, K_d, K_p	= transport coefficients of convection mixing, vertical concentration back diffusion and parasitic mixing, respectively
L	= effective length of the column
m_t, m_p, m_+, m_-	= mass of gas in the total sample, in parasitic volumes and in the positive (enrichment) and negative (storage) reservoir, respectively
m_+''	= fictitious mass in positive reservoir obtained by correcting m_+ for enrichment in the part of the column adjacent to the positive reservoir and depletion in the part adjacent to the negative one
M	= molecular weight
n	= number of columns in parallel
p	= pressure
q	= separation factor, <i>ie</i> , the ratio of the isotope ratios in the enrichment reservoir and in the storage reservoir
q'	= enrichment, <i>ie</i> , the ratio of the concentrations of the enriched isotope in the enrichment reservoir and in the original gas
q_e, q_e'	= equilibrium, <i>ie</i> , maximum, value of the separation factor and the enrichment, respectively
r_1, r_2	= radii of outer and inner tube (or hot wire), respectively
t	= time
t_r	= relaxation time, <i>ie</i> , the time in which the thermal diffusion system would reach its equilibrium state if the initial transport persisted
T	= absolute temperature of the gas (K) $T_1 \leq T \leq T_2$, where subscripts 1 and 2 refer to cold and hot wall, respectively
V_+, V_-	= volume of the positive bottom and the negative top reservoir, respectively
z	= coordinate along the column ($0 \leq z \leq L$) with positive direction towards the positive reservoir
α	= thermal diffusion constant
μ	= mass of gas per unit length of column
τ_1, τ_2	= transport along the column of species 1 and species 2, respectively (in mass per unit time)

and (b) the performance of the enrichment system in terms of the initial transport and the following transient phase.

Single column

Thermal diffusion (TD) columns have been discussed by Jones and Furry (1946) and by Vasaru *et al* (1969). In this paper, the notation of Jones and Furry will be used mostly. A list of symbols and definitions is given in table 1.

The TD column consists of a vertical tube which is cooled on the outside with a heated wire along its axis. In a horizontal cross section, the temperature gradient effects a small isotope separation in which, generally, the heavy isotopic molecules concentrate near the cold wall and the light ones near the hot wire. Convection currents, going down at the cold wall and up in the warm center part of the column, then result in a net transport of the heavy isotopes down the column and of the light isotopes to the top. As a result, a concentration gradient is built up along the tube.

The net transport of isotopes through the column is determined by 4 processes (fig 3):

- (i) the TD process, which builds up a horizontal concentration gradient;
- (ii) a horizontal back diffusion (HBD) due to this concentration gradient;
- (iii) a convection current (CC), transporting the isotopically heavy gas downwards along the cold wall and the light gas upwards along the wire;
- (iv) a vertical back diffusion (VBD) due to the vertical concentration gradient.

The convection current (CC) provides the net transport of isotopic molecules. However, when a vertical concentration gradient has been established, the current also replaces, in a horizontal cross section, the gas near the wire by isotopically heavier gas from below and near the cold wall by isotopically lighter gas from above. This reduces the concentration difference between the hot and the cold gas flow.

The net mass transport τ of a certain, rare isotope in a TD column is given by (Vasaru *et al*, 1969)

$$\tau = HC - (K_c + K_d) \partial C / \partial z \quad (1)$$

where H, K_c and K_d are system parameters called transport coefficients and C is the concentration of the isotope considered ($C \ll 1$). It is the difference between the thermal diffusion transport HC (H proportional to p^2 and r_1^4) and the convection mixing $K_c \partial C / \partial z$ (K_c proportional to p^4 and r_1^8) and diffusion mixing $K_d \partial C / \partial z$ (K_d proportional to r_1^2). Here, p is the gas pressure in the column and r_1 is the radius of the hot wire column. Thus, a high thermal diffusion transport may be obtained in a large diameter tube at high pressure. This results only in a large net mass transport in the initial phase when concentration gradients are small and, consequently, diffusion and convection mixing are negligible. In the fol-

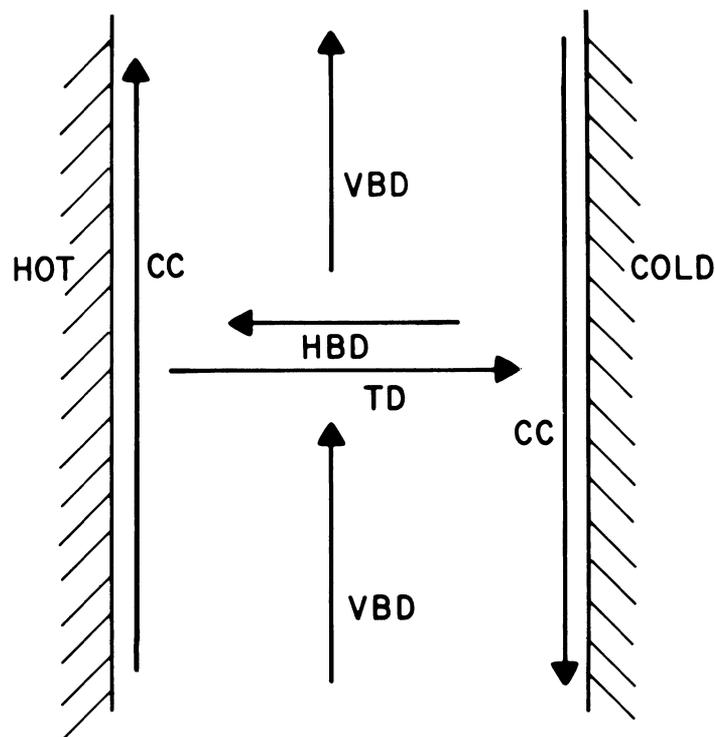


Fig. 3. Processes determining the net transport of isotopes in a thermal diffusion column; we consider the transport of the heavy isotopic species. TD—thermal diffusion isotope separation, HBD—horizontal back diffusion, CC—convection current, VBD—vertical back diffusion.

lowing, transient phase net mass transport gradually decreases to zero as a concentration gradient along the column is built up. The decrease is determined by the approach of the concentration distribution to its equilibrium state where TD transport is balanced by convection mixing and concentration back diffusion. To describe the isotope separation, we define the separation factor q (\equiv isotope ratio on the enrichment side/isotope ratio on the depletion side of the column) and the enrichment q' (\equiv the concentration of the (heavy) isotope considered in a certain volume divided by the original concentration of this isotope in the gas).

To obtain a net mass transport that persists in the transient phase until the desired enrichment in the enrichment reservoir has been obtained, we must make the maximum ($=$ equilibrium) separation factor q_e much larger than the q at which enrichment will be stopped. With $\tau = 0$ in equation (1) we derive:

$$q_e = \exp (HL / (K_c + K_d)), \quad (2)$$

where L is the effective length of the column.

Because of the different functional dependences of H , K_c and K_d on column radius r_1 and pressure p a maximum of $q_e(p)$ is obtained if $K_d = K_c$ and of $q_e(r_1)$ if $K_d = 2K_c$. A practical system will result from a compromise between a high initial net mass transport (large H , but also $K_c \gg K_d$) and a maximum separation factor. Enrichment columns 0.9 to 1.0cm in diameter filled with 800 to 900 torr of CO, having a cold wall temperature of about 300 K and a 0.3mm diameter wire at about 950 K are such a practical compromise (K_c is 1 to 2 K_d) (cf Grootes, 1977).

Figure 4 shows a theoretical calculation of q_e for $^{12}\text{C}^{18}\text{O}$ as a function of pressure for different hot wire temperatures T_2 for a column of 1.0cm diameter (top section). At low pressures diffusion mixing (independent of p) dominates over convection mixing and q_e increases with pressure. For high pressures, however, the convection is dominant and, consequently, there q_e decreases with increasing pressure.

The enrichment system

Both enrichment, and enrichment time, depend not only on the columns used, but also on the size of the reservoirs and on the connections between the columns and between columns and reservoirs.

The separation factor of a system of 2 or more columns in series is simply the product of the separation factors of the individual columns. The enrichment is determined mainly by the relative size of the different

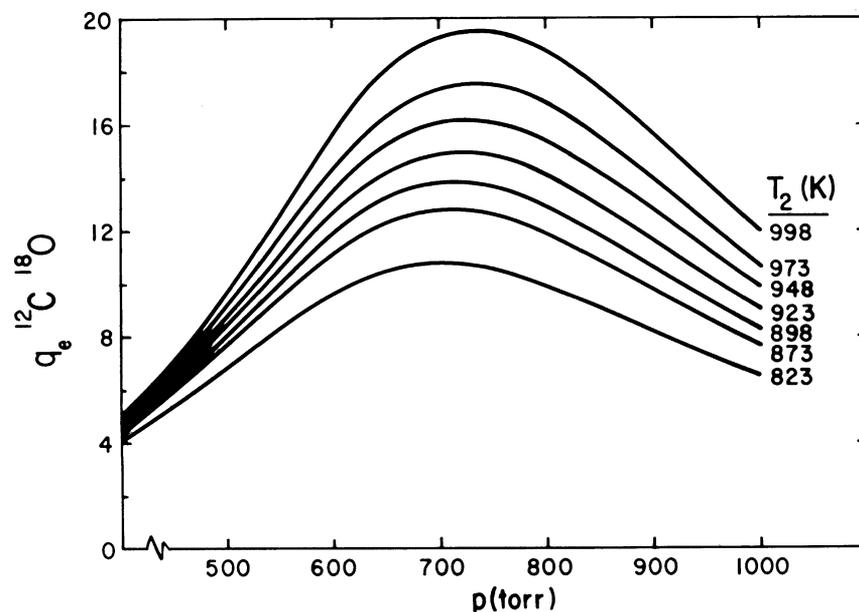


Fig. 4. Theoretical separation factor q_e for $^{12}\text{C}^{18}\text{O}$ of a single top column ($r_1 = 0.5\text{cm}$) as a function of pressure and wire temperature at a cold wall temperature $T_1 = 300\text{K}$.

parts of the enrichment system over which the finite amount of the heavy isotopic species present in the total sample of mass m_t is distributed, provided the equilibrium separation factor $q_e \gg m_t/m_+$, *eg*, $q_e > 10 m_t/m_+$, where m_+ is the mass of the gas in the enrichment volume V_+ . This means that although q_e of a system may be quite high (> 1000) the maximum enrichment may still be small if m_t/m_+ is small. Because the initial net mass transport of an enrichment system is determined by the transport properties of the individual columns, and their combination, it is independent of the mass ratio m_t/m_+ . If we compare the enrichment in systems having largely different values of m_t and the same m_+ , this means that initially both systems will give the same enrichment as a function of time. When enrichment in the small m_t system approaches q'_e , however, the enrichment in the large m_t system will continue to increase till it also approaches its much larger equilibrium value after a proportionally longer time. To determine the enrichment of a system we thus have to consider it in detail.

System geometry

Figure 5 shows an enrichment system of n parallel columns with $r_1 = 0.50$ cm, each with its own storage reservoir, in series with a single column with $r_1 = 0.45$ cm. The TD transport of n parallel columns is nHC . This means that if concentration C increases along the columns the number of columns can be decreased to give the same transport. Thus the use of two stages with the larger number of columns in parallel in the first is favorable. In the next section we discuss which value of n gives the largest TD transport for a given number of TD columns, *ie*, a given input in construction and operating costs.

The driving force of the isotope separation is the TD transport in the top and bottom section columns. Concentration gradients in the system reduce the net mass transport. In the first place, this is the inevitable, gradually increasing, concentration gradient caused by the enrichment in V_+ . Secondly, in three regions of the system, transport must take place by concentration diffusion. These are from V_- (*ie*, the convection flow of reservoir gas through the top of the column) to the beginning of TD transport (*cf* Rutherford and Kaminski, 1967), from the bottom of the top section to the top of the bottom column and from the bottom of this column to V_+ . Concentration gradients opposite to the ones in the columns exist in these regions. In the third place the concentration at the top of the bottom column has to be considerably higher (about n times, $\tau = HC - K\partial C/\partial z$) than that at the top of the n top columns. The effect of these restraints on the concentration distribution in the system will be to reduce the net mass transport of the system even before any appreciable enrichment in V_+ has taken place.

If accumulation of heavy isotopes in and between the columns is small compared with the total transport, which it is, we may neglect it and assume a constant transport τ through the system, quasi-stationary assumption. Under this assumption the TD transport of a single column

after a time t has been given by Felber and Pak (1974). Combination of their TD mass transport equation for the columns with expressions for the concentration diffusion mass transport between the columns and between columns and reservoirs gives a set of equations that, with some simplifying assumptions, may be solved. In figure 6, we compare the calculated initial mass transport of a combination of n columns to 1 with that obtained if we take the same number of columns combined 1 to 1. For $n < 6$ the combination n to 1 is more efficient than 1 to 1. The best result, an increase in TD transport with about 15 percent, is obtained for 3 to 1.

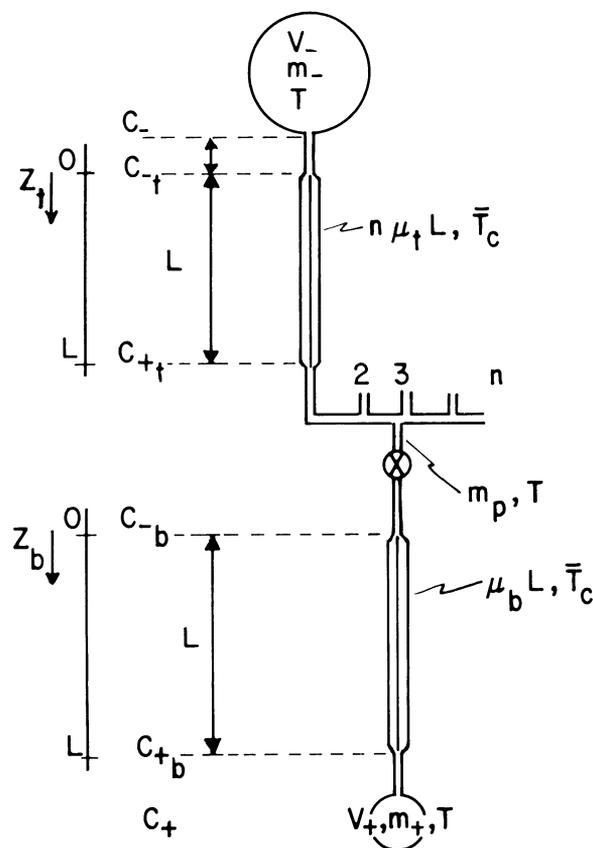


Fig. 5. Enrichment system of n parallel top columns, each of effective length L and with its own storage reservoir V_- , in series with a single bottom column of the same length. V_+ —enrichment reservoir; m_+ , m_- , m_p —mass of gas in V_+ and V_- and parasitic mass in connecting tubes, respectively; $n\mu_t L$, $\mu_b L$ —mass in the n top columns and in the bottom column resp., where μ is the mass per cm in the columns; T , \bar{T}_c —room temperature and average temperature of the gas in the columns; C_- , C_{-t} , C_{+t} —concentration of the heavy isotope in V_- , at the top and bottom of the top section; C_{-b} , C_{+b} , C_+ —concentration of the heavy isotope at the top and bottom of the column and in V_+ .

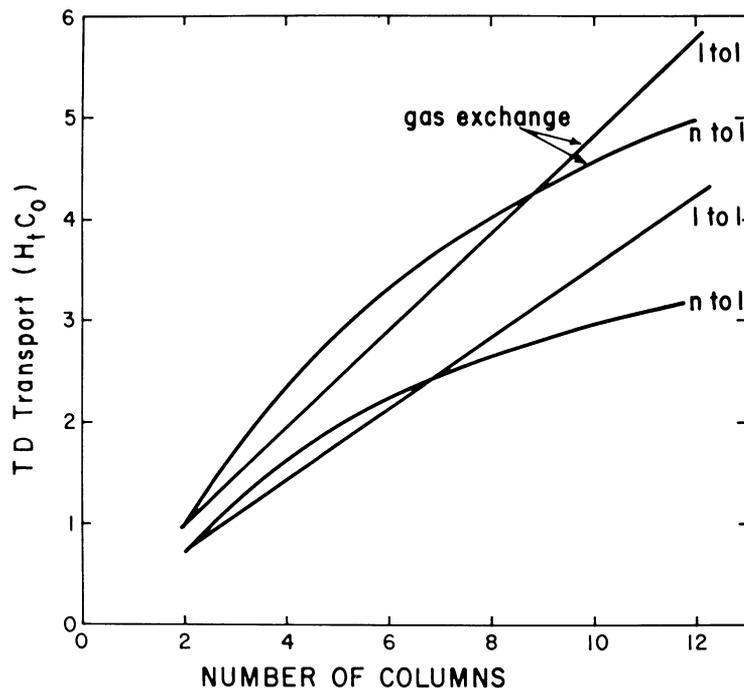


Fig. 6. Thermal diffusion transport in units $H_t C_o$ for $n+1$ columns combined 1 to 1 in series and combined n parallel with 1 in series if no concentration diffusion gradients are present (gas exchange) and if concentration diffusion does exist (no gas exchange). The indices t and o refer to top column and the original concentration, respectively.

If we eliminate the necessity of concentration diffusion transport by forced gas exchange between the bottom of the top section and the top of the bottom column, and between V_- and the top of the top section and V_+ and the bottom of the bottom column, we can considerably improve the TD transport of the system (fig 6, gas exchange). For this situation $n \leq 7$ is more efficient and the best result, about 22 percent gain, is obtained for 3 to 1. The UW system consisting of 3 units of 4 to 1 and 2 of 3 to 1 (fig. 1) therefore has near optimum geometry.

The equilibrium enrichment

The equilibrium enrichment q'_e is found from q_e and a mass balance of the system. Figure 7 shows q_e and q'_e calculated for a 4 to 1 system at 850 torr pressure as a function of the hot wire temperature T_h for 2 different cold wall temperatures. The mass ratio m_t/m_+ ($= 16.9$ for this system) obviously is the main factor determining q'_e in this case. A positive effect of a dominance of m_t/m_+ is the attenuation of the effect of variations in hot and cold wall temperature. Their relative effect on q'_e is 10 to 20 times less than on q_e . Thus, the fluctuations in the cooling water temperature (cold wall temperature range 25-45°C) will only produce

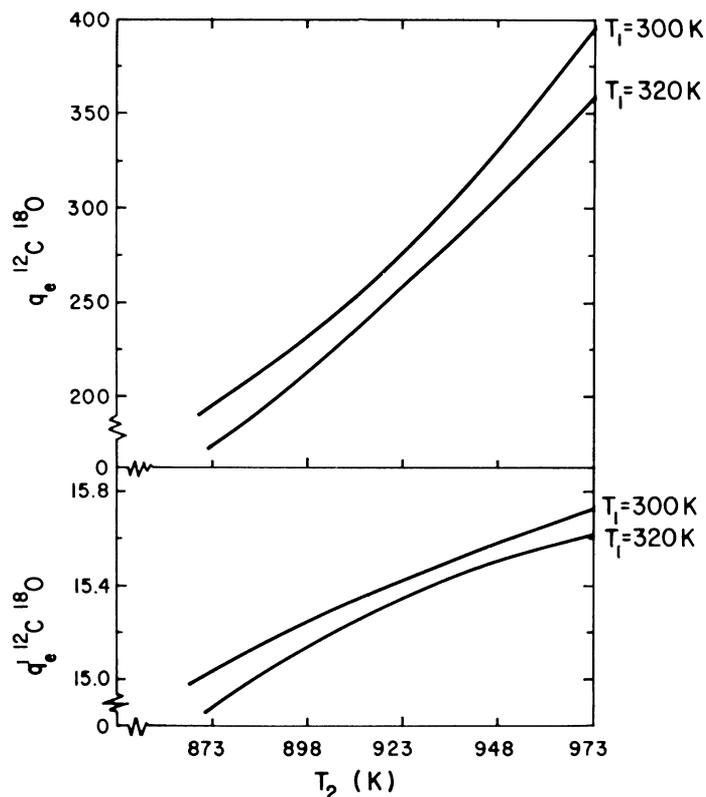


Fig. 7. The equilibrium separation factor q_e and the equilibrium enrichment q'_e as a function of wire temperature T_2 for $^{12}\text{C}^{18}\text{O}$ in a 4 to 1 system ($r_{1\text{top}} = 0.5\text{ cm}$, $r_{1\text{bottom}} = 0.45\text{ cm}$) at 850 torr and $T_1 = 300$ and 320 K.

variations in q'_e of 0.5 to 1%. The enrichment of the system is therefore rather reproducible and stable. The change in calculated equilibrium enrichment q'_e with the mass ratio m_t/m_+ is shown in figure 8 for a 4 to 1 system with $q_e = 276$. The different mass ratios of the UW system (a) and that routinely used in Groningen (b) may be the main reason for the difference in enrichments. When $q_e > 10 m_t/m_+$ the difference between q'_e and m_t/m_+ is $< 13\%$ and q'_e increases linearly with m_t/m_+ . Only when $q_e < 10 m_t/m_+$ (small q_e , large m_t/m_+) this linear relation breaks down. In practice, the equilibrium enrichment is never reached because, in the final stage, the increase of enrichment with time is very slow. Therefore the transport and enrichment of the system in the initial and the transient phase are of practical interest.

The initial phase

The initial phase lasts only as long as no significant depletion or concentration build-up occurs. Since quasi-stationary transport in our system requires considerably different concentrations in different parts of the

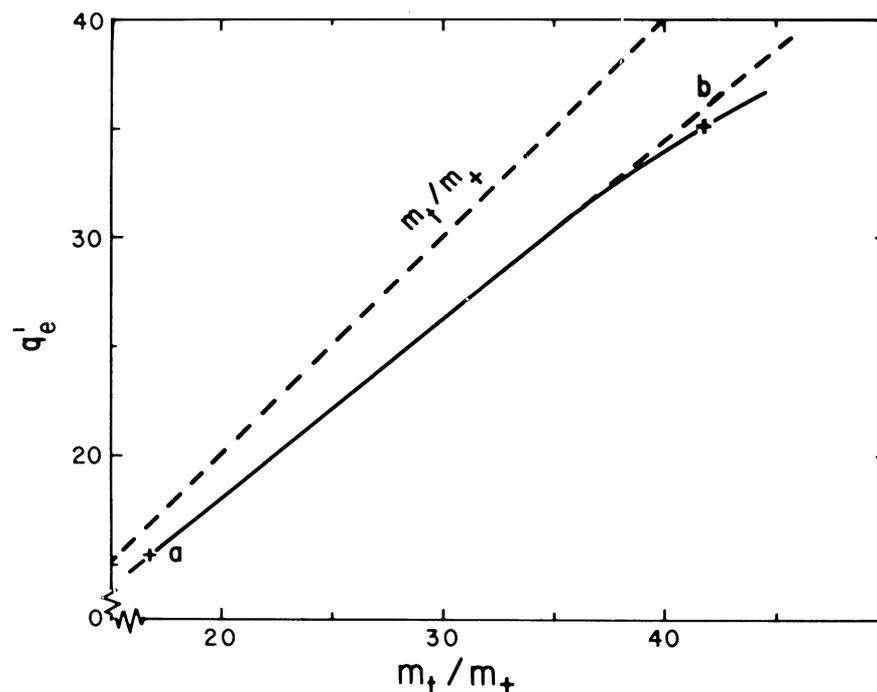


Fig. 8. Relation between the equilibrium enrichment factor q_e' and the ratio total sample mass/mass of the enriched sample, m_t/m_+ , for a 4 to 1 system with $q_e = 276$. a—the UW system, b—the routine value of the Groningen system.

system (even with gas exchange), this phase will last for a very short time, probably less than a day. For 7 enrichments under varying conditions, we measured the enrichment as a function of time by taking small samples of the gas in V_+ .

The first was at lower wire temperature (≈ 860 K) without gas exchange as was routine before July 1977. The second was a trial in which 2 systems were subjected to forced gas exchange between top and bottom columns and between columns and reservoirs and 2 were heated to a higher temperature (≈ 960 K). The other 5 combined the high temperature with different gas exchange frequencies. Gas exchange was effected by a gas swing, heating the top reservoirs V_- with heating tape for about 30 min, which caused a pressure increase of about 50 torr, and cooling for another 30 min. The amount of gas forced out of V_- is a little more than that needed to replace the gas in the connecting tubes.

The TD transport coefficient H can be found from the initial enrichment. From eq 1, it follows, for a single column, that in the initial phase, before concentration gradients have been built up, $\tau = HC$. Measuring the increase in concentration as a function of time enables us to calculate the transport, τ . From this transport during the first few days of the en-

richment, and the isotopic composition of the gas (C), we obtain an experimental measure of the TD transport coefficient of the system. In a composite system like ours transport is not only determined by the columns, but also by their combination. The calculation, therefore, yields an effective system transport coefficient H_s . The observed H_s is a minimum value for the system TD transport coefficient, because we can only determine it from the build-up of a concentration gradient in the system. Nevertheless, it allows an estimate of the agreement between the calculated TD transport coefficient of the columns and reality.

To get sufficient measurements the results of the first 4 days were used to determine the initial transport. We calculated H_s from

$$H_s = \frac{M_2}{M_1} m'_+ \frac{dq'}{dt}. \quad (3)$$

Here M_2 and M_1 are the molecular weight of the heavy and the light isotopic species, respectively and m'_+ is the mass in the enrichment reservoir plus that in the lower column. The factor M_2/M_1 is necessary because H_s gives the mass transport of the heavy isotopic species, while m'_+ is based on the actual gas, mainly the light isotope. The use of m'_+ instead of m_+ reflects the considerable concentration build-up in the bottom column during the first phase of the enrichment necessary to obtain a constant transport through the system. The experimental system TD transport coefficients H_s were divided by the number of columns of the system (4 or 5) to get a measure of the net separating mass transport obtained per unit investment (column, power, cooling water) (table 2). This also provides a basis for comparison with the Groningen system.

From table 2 it seems that increasing the wire temperature from 860 K to 960K (average gain 24 percent) was less important than reducing the concentration diffusion gradients by gas exchange (average gain 57 percent). The observed overall increase in initial transport of about 80 percent is considerable and results in a much faster approach to the equilibrium enrichment.

To compare the observed system TD transport with theory, we first calculated the TD transport coefficients H for the different experiments from the gas properties, system parameters, and the experimental TD coefficient α (Boersma-Klein and de Vries, 1966) using the Lennard-Jones 12-6 potential model. From H we calculated system TD transport coefficients H_s taking into account the reduction in transport caused by the gradients of concentration diffusion in connecting tubes and/or by the higher concentration at the top of the bottom column required by the n to 1 system geometry (calculation to be published elsewhere). Dividing H_s by the number of columns in the system allows a comparison with our experimental results (table 2). Observed transports are only 50 to 60 percent of those calculated. The calculated reduction of the initial system mass transport due to concentration gradients depends on both H and K . The presence of parasitic mixing (K_p , not considered in the calculation) and of factors reducing H and/or increasing K will, therefore, lead to an

observed mass transport that is lower than calculated. Also the use of measurements over the first 4 days yields a minimum value for the initial transport and may account for part of the observed discrepancy. The agreement between calculated and observed gain is excellent.

The calculated mass transports for the 5 systems at 960 K, with gas exchange, are rather similar. We can, therefore, use the observed transports in table 2 to compare the relative system performances. The difference in transport between supposedly identical systems is of the order of 20 percent for #1 and #5. The difference between #1 and the three 4 to 1 systems is approximately 10 percent and equals one standard deviation. This means that these 4 systems are not significantly different. Only #5 is less efficient. The difference of 20 percent is of the same order as the maximum difference observed between the single 1 to 1 columns in Groningen (Grootes *et al*, 1975) and apparently indicates the variation encountered in the construction of “identical” hot-wire column systems. The initial TD transport of Groningen column 1 (performance close to the average of all 9 columns) gives 1.79×10^{-6} g/sec/column or 16 percent less than the average value of 2.13×10^{-6} for the 5 UW systems. The Groningen system was not subjected to gas exchange and operated at ≈ 920 K. Since a calculation indicates that the concentration diffusion gradients reduce the initial transport by about 20 percent, we may conclude that the physical properties of the UW and the Groningen systems are very similar as far as the initial transport is concerned.

TABLE 2
Initial mass transport for the UW systems under different
operating conditions

		Transport mass 30 (10^{-8} g/sec/column)			
		Original $T_2 = 860$ K	Gas exchange $T_2 = 860$ K	$T_2 = 960$ K	Gas exchange $T_2 = 960$ K
#1	Observed	1.54	2.02		2.39 ± 0.16^a
	Calculated	2.18	3.28		3.99
#2	Observed				2.13 ± 0.21
	Calculated				3.98
#3	Observed		1.46		2.06 ± 0.13
	Calculated		3.16		3.92
#4	Observed	1.04		1.17	2.13 ± 0.14
	Calculated	2.08		2.55	3.93
#5	Observed			1.24	1.94 ± 0.12
	Calculated			2.62	4.03
Average gain	Observed		57%	24%	80%
	Calculated		53%	23%	86%

^a uncertainty in mean, based on observed scatter of measurements from different experiments; variability of transport due to small changes in TD parameters like column temperature and filling pressure is large, this contributes to overall uncertainty given; within each single experiment the difference between the different systems is present with a smaller standard deviation.

The transient phase

In figure 9A, B, we give the enrichment in system 1 and 4, respectively, as a function of time for 5 enrichments. When we compare the enrichment under the original conditions (•: $T_2 = 860$ K, no gas exchange) with the others, we see that the main improvement of the new conditions is a faster approach to equilibrium. Between 10 and 30 days the enrichment is significantly higher than before, but after 40 to 50 days the change is small. The results for the enrichment during which gas exchange was increased (x: coal-9, 19 min heating/cooling, pressure difference ≈ 60 torr) are similar to the other enrichments with gas exchange during the first 2 weeks but then fall below the others. After about 4 weeks equilibrium seems to have been reached at a value of q' about 15 percent (#1) and ≈ 30 percent (#4) below that of the other enrichments. This indicates that forced gas exchange is not just beneficial in that it reduces concentration diffusion gradients and thus enhances TD transport, but that it probably also results in some mixing in the columns and disturbance of the concentration gradient built up in the columns by TD. For this reason, we henceforth stopped the forced gas exchange after ≈ 17 days. By that time, TD transport decreased considerably due to depletion of V_- and the build-up of a concentration gradient over the column so the disturbing influence becomes more important than the gain due to the reduction of concentration diffusion.

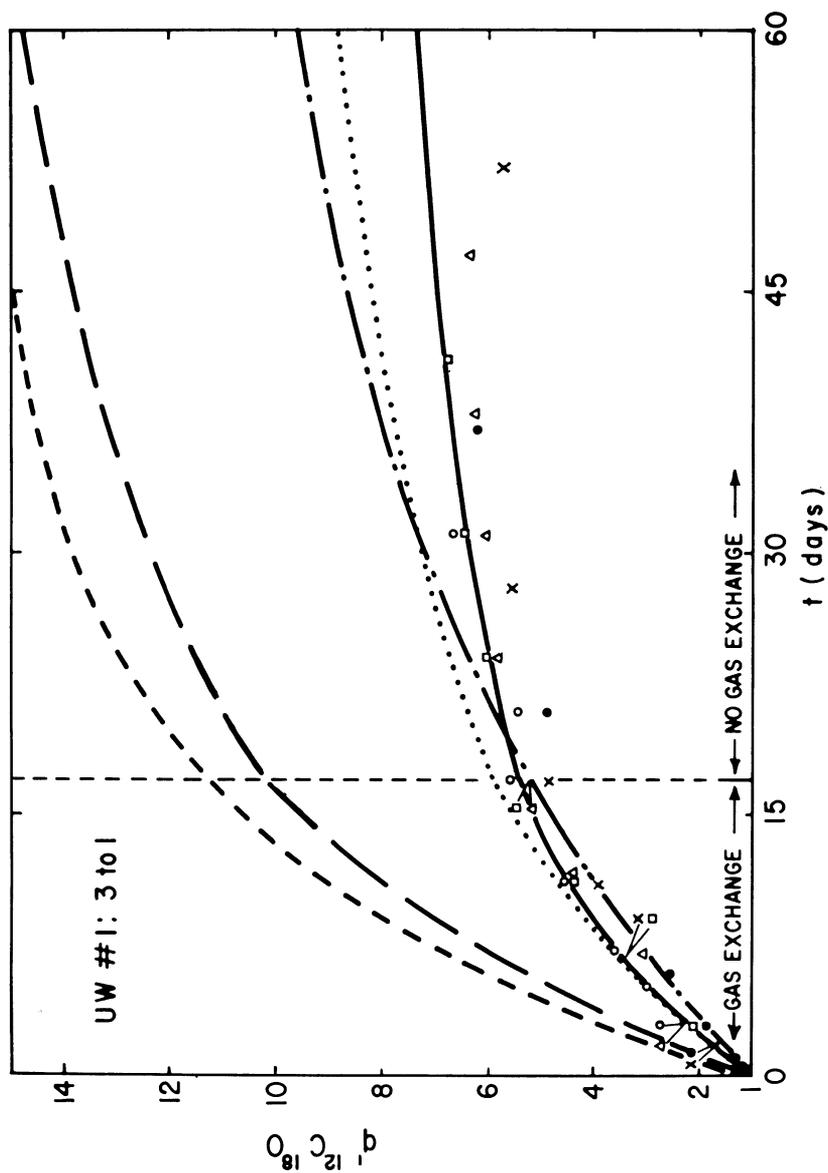
Although the 3 other enrichments shown in figure 9A,B were all under somewhat different conditions, the difference in enrichment at a certain time generally is of the order of 10 percent or less. The enrichment systems seem to work rather reproducibly. The attenuation of the influence of variation of parameters like hot and cold wall temperature and pressure by the constant ratio m_-/m_+ probably is an important factor in this system reproducibility. It enables us to predict with reasonable accuracy the enrichment obtained after a certain time and thus to choose an enrichment time according to our needs.

The enrichment during the transient phase was calculated using an approximation described by Felber and Pak (1974) in which enrichment and depletion of the gas in the columns are accounted for by adding a correction term to the mass of the gas in the enrichment reservoir, while the columns are treated as having a volume zero. This way the enrichment q' as a function of time t can be approximated by a simple exponential:

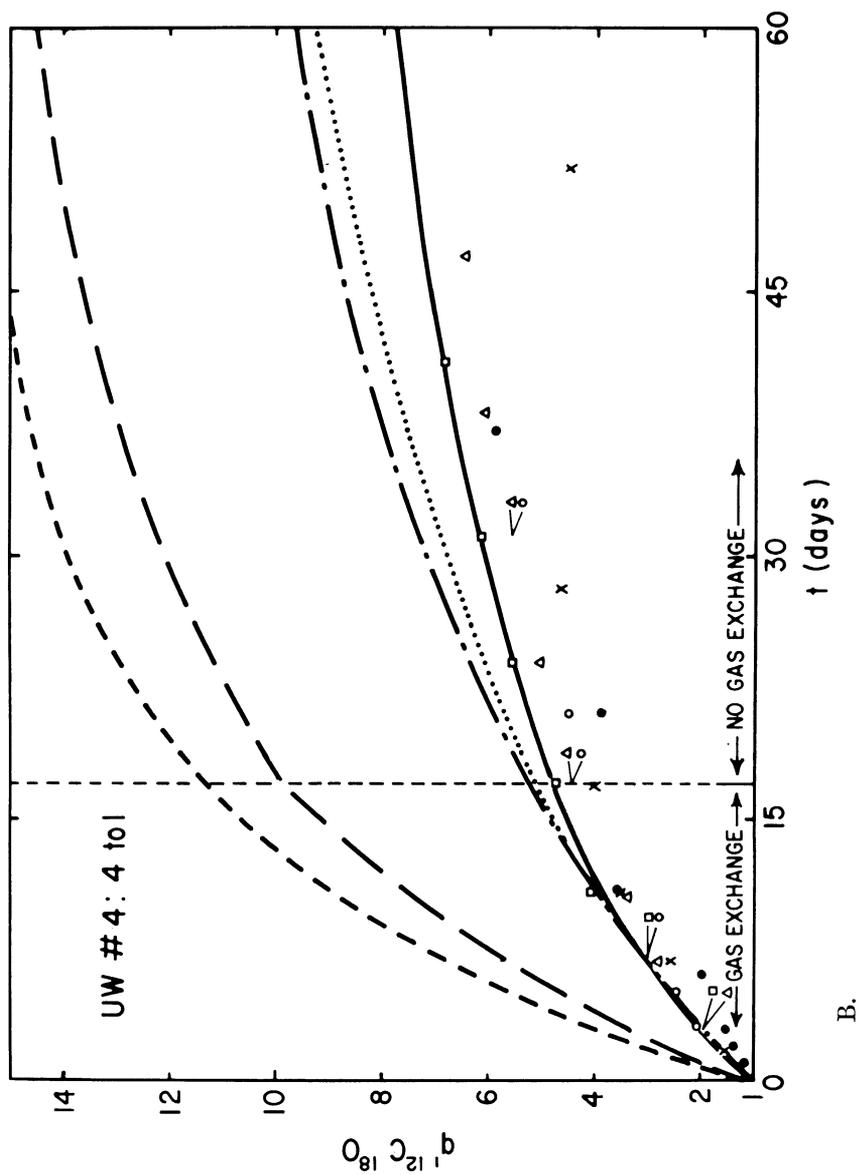
$$q'(t) = q_e' - (q_e' - 1)\exp(-t/t_r), \quad (4)$$

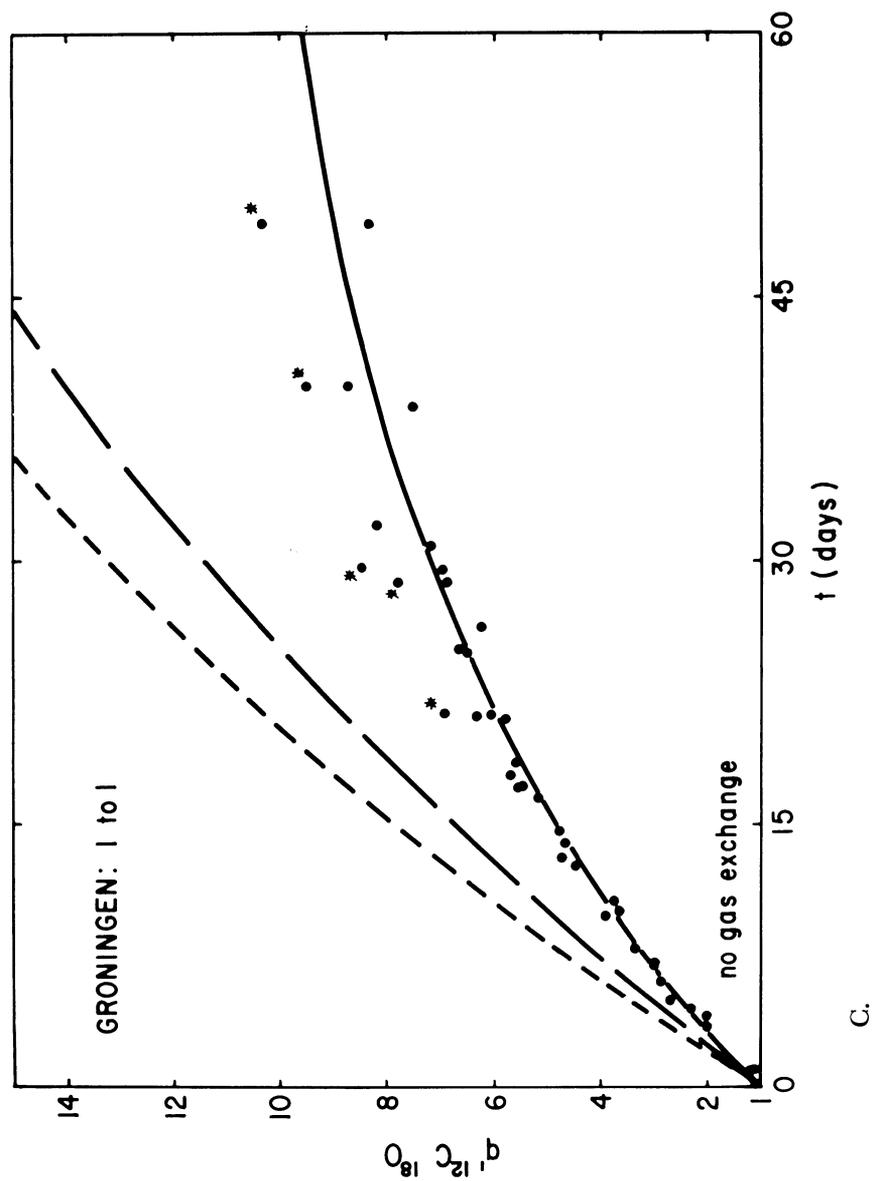
where q_e' is the final equilibrium value of q' and t_r is the time in which q_e' would be reached if the initial transport continued. To give a realistic description of the enrichment system the constraints on the mass transport, due to the connection of 3 or 4 columns to 1 and to the concentration diffusion mass transport through connecting tubes, have to be considered. Using eq 4, we calculated the increase in enrichment during consecutive short time intervals for quasi-stationary transport with these con-

Fig. 9A,B,C. Enrichment of mass $30/^{12}\text{C}^{18}\text{O}$ as a function of time; experimental points: mass 30 ($^{12}\text{C}^{18}\text{O}$, $^{13}\text{C}^{17}\text{O}$, $^{14}\text{C}^{16}\text{O}$), calculated lines: $^{12}\text{C}^{18}\text{O}$; o—marble M-23, 48 min in gas exchange cycle; x—coal 9, 38 min gas exchange cycle; Δ —Q-L-114, 64 min gas exchange cycle; \square —Q-L-1210, 64 min gas exchange cycle; \bullet —marble M-21, no gas exchange. Marble M-21 enriched with 55 V wire heating, (860 K). ——— best fit of H and K to experimental points; ——— calculated enrichment for actual system geometry and conditions; $\cdots\cdots\cdots$ calculated enrichment for idealized system; $-\bullet-\bullet-$ best fit. Groningen results for comparison in UW graph; $\cdots\cdots\cdots$ calculated enrichment for doubling V_{∞} ; $\cdots\cdots\cdots$ calculated enrichment using the effective H and K values of the best fit.



A.





straints and the initial and the theoretical equilibrium state as boundary conditions (fig 9A,B). The calculations will be described elsewhere.

Because gas exchange is stopped after 17 days, the predicted enrichment curve is that for a sample with ideal gas exchange for the first 17 days and, after that, for a system without gas exchange continuing from the enrichment built up in those first 17 days. Comparison of the increase in enrichment when the enrichment system geometry is considered in detail, with that of an “idealized” system, shows that the more accurate description predicts 10 to 20 percent lower enrichments in the transient phase. Observed enrichments, however, are still considerably lower (by as much as 50 percent). To evaluate this difference in terms of our transport coefficients H and K ($= K_c + K_d (+K_p)$), where K_p represents parasitic mixing, we varied the values of H and K used in our calculation till a satisfactory fit to experimental results was obtained. Because the enrichments in figure 9A,B were all under slightly different conditions, we chose one, □: QL – 1210, to make our fit. This run gave the best enrichment and was considered most representative of present system operation. Because the gas exchange probably also introduces some parasitic mixing in the columns, which would be apparent as a larger K_p , we made separate fits for the first 17 days and the remaining enrichment, keeping H constant. The results are given in table 3.

An independent estimate of both H and K is obtained, because initially the enrichment is mainly determined by H ; but near equilibrium by H and K combined. The value of 0.8 H (plus or minus a few percent) giving the best fit in both geometries combined with the larger K would lower the calculated initial mass transports in table 2 by about 40 percent. The observed values, then, are about 90 percent of those calculated. Since over the first 4 days not the true initial transport, but rather the first part of the transient phase was sampled, a lower experimental H can be expected. The lower K , when gas exchange stops, gives a quantitative indication of the disturbance and parasitic mixing introduced by gas exchange. For the 4 to 1 system considered it is apparently more serious than for the 3 to 1. The values of H and K are very sensitive to the cold wall temperature T_1 and to the column radius r_1 . Since H , K_c and K_d are influenced by these parameters in the same direction, however, they cannot explain

TABLE 3
Effective transport coefficients deduced from experimental enrichment results

System	Condition	$H_{\text{eff}}/H_{\text{theory}}$	$K_{\text{eff}}/K_{\text{theory}}$
#1	Gas exchange, $t < 17$ days	0.8	1.85
3 to 1	No gas exchange, $t > 17$ days	0.8	1.75
#4	Gas exchange, $t < 17$ days	0.8	2.0
4 to 1	No gas exchange, $t > 17$ days	0.8	1.6
Groningen	No gas exchange	0.95	2.0

both the observed discrepancies. A value for the TD constant α (in H) that is 20 percent too high and a very large K_p (0.6 to 0.7 ($K_c + K_d$)) act independently and would explain the observations. Variations in r_1 and T_1 along the column (the latter due to the temperature increase of the cooling water during its contact with the column) affecting H and K differently may also account for part of the observed discrepancies.

In figure 9C, 5 enrichments with the Groningen TD system are given (Grootes, 1977). Using the same calculating procedure as for 9A,B, a line was fitted to the points. These mostly resulted from samples taken from column #1 for $V_- = 201$ L ($3\times$) and $V_- = 231$ L ($2\times$), but the compositions of the final, averaged enriched samples of all 9 columns are also given (*). The line gives an average performance of column #1 for the 2 closely similar situations. If sampling disturbs the enrichment, this will provide the best comparison with the UW systems of figure 9A,B which were likewise regularly sampled. It should be noted, however, that for these enrichments the average enrichment of the complete system is 10 to 15 percent higher than that of column #1. For comparison the line fitted to the experimental points of the Groningen column #1 is given in figure 9A,B.

The enrichment of the Groningen system is closely similar to that of UW system #4 in figure 9B and somewhat lower than that of #1 in figure 9A during the first 2 weeks. Since its rate of enrichment decreases less rapidly, the Groningen enrichment becomes larger than that of the UW systems after about 3 weeks. The effective values of H and K that fit the experimental results deviate from theory in a similar way as those of the UW systems (table 3), although they are relatively somewhat larger. This agrees with the similar results obtained for H in the initial phase. To evaluate the influence of sample size on the enrichment, we calculated the enrichment of the UW systems as a function of time for a doubling of the reservoir size V_- (fig 9A,B) and for a V_- giving the same m_t/m_+ as in Groningen, using the effective values of H and K from table 3. The same m_t/m_+ resulted in an enrichment curve (not shown) that closely approached that of the Groningen system after 4 to 5 weeks. This suggests that the two enrichment set-ups, although of different geometry, behave similarly not only in the initial but also in the transient phase and that the difference in observed enrichment is mainly caused by the different ratio m_t/m_+ . From figure 9A,B we see that doubling V_- will increase the enrichment by about 10 percent after 30 days and about 15 percent after 40 days. This corresponds to a gain in age range of 800 to 1100 years. Therefore, although some gain in age range can be obtained by increasing the sample size, efforts to reduce K may, for the moment, be more effective. Here it is important to note, that, although the UW systems have been constructed without spacers to center the wire, thus introducing possible parasitic convection mixing, their effective K values are somewhat lower than that of the Groningen system in which the wire is centered by triangular spacers. The cause of the high K values observed therefore deserves further attention.

Figure 10 show the enrichment of all 5 UW systems during a typical enrichment under the new conditions, *ie*, $T_2 = 960$ K and gas exchange during the first 17 days. The differences between the individual systems found for the initial transport (table 2) obviously persist through the transient phase. The enrichment can best be stopped after about 4 weeks, for after this time the increase in enrichment slows. In 4 weeks with the present system, a sample enriched by a factor ≈ 5.5 in ^{18}O or about 7 in ^{14}C will be obtained giving a gain in age range of almost 16,000 years. The actual ^{14}C enrichment of a sample is calculated from the mass spectrometrically measured enrichment of mass 30 (mainly $^{12}\text{C}^{18}\text{O}$) and an experimentally established relation between this enrichment and that of ^{14}C . The uncertainty in the ^{14}C enrichment thus obtained is a few percent ($<5\%$). This is small compared with the statistical uncertainty in the ^{14}C activity of old samples and generally has a negligible influence on the uncertainty in the measured sample age.

CONCLUSIONS

(i) The combination of 3 or 4 columns in the top section to 1 column in the bottom section offers an advantage of about 20 percent for a system with gas exchange or 15 percent for a system without over a simple 1 to 1 system as far as initial thermal diffusion transport for a given energy consumption is concerned. This is the optimum value for the UW system.

(ii) The initial transport of the systems was increased by as much as 80 percent by increasing the hot wire temperature from 860 to 960 K ($\approx 24\%$) and by gas exchange ($\approx 57\%$).

(iii) The observed initial net thermal diffusion mass transport for mass 30 is of the order of 2.1×10^{-6} g/sec/column and 16 percent larger than in Groningen (1.8×10^{-6} g/sec/column). The difference is due to different operating conditions.

(iv) The mass ratio m_t/m_+ is an important factor determining the enrichment provided $q_e \gg m_t/m_+$.

(v) The increase in initial transport due to gas exchange and temperature increase of the wire results in a faster approach to the equilibrium enrichment without substantially increasing this value.

(vi) The enrichment of the UW system in the transient phase is similar to that of the Groningen system. Differences are mainly attributable to the difference in m_t/m_+ .

(vii) The similarity in the enrichments of the UW and the Groningen systems suggests that hot wire column enrichment system geometry is not very critical, and that it is relatively easy to construct a system that works reproducibly and satisfactorily.

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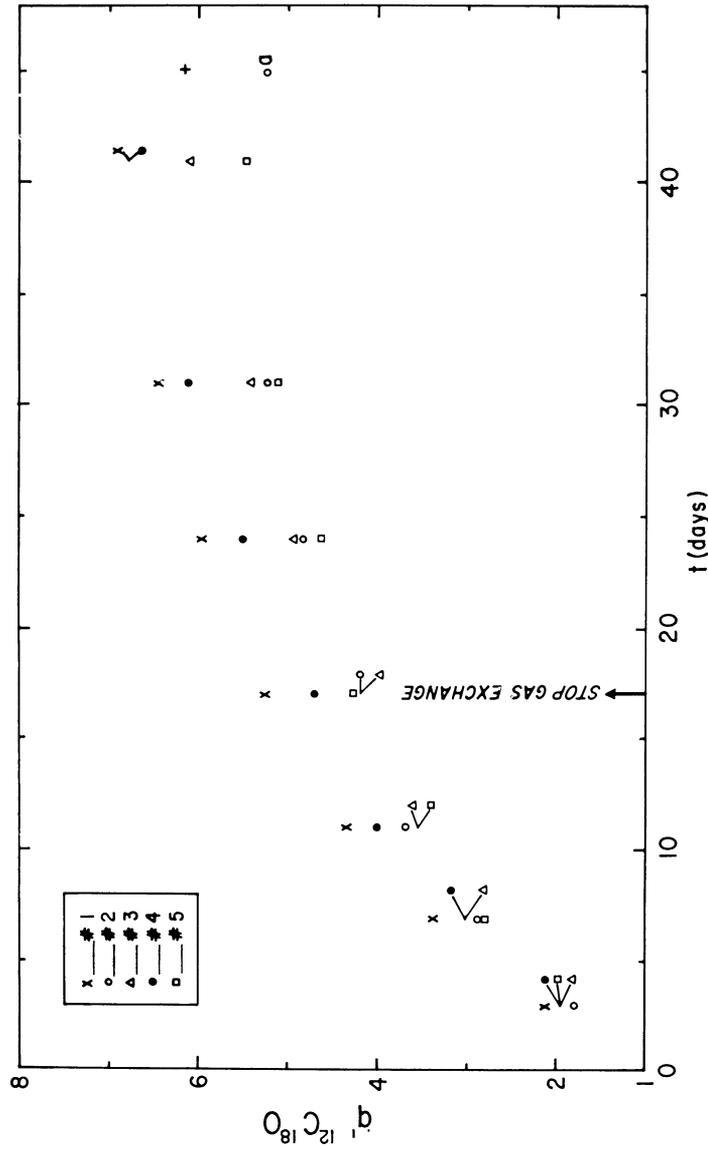


Fig. 10. Enrichment of $^{12}\text{C}^{18}\text{O}$ as a function of time for the 5 UW systems. After 17 days the gas exchange was stopped. a—final sample of system #2 shows too low enrichment because of a small leak at the bottom of the column; +—final enriched sample, mixture of all 5 systems.

present study were taken. P J Wilkinson took care of sample pretreatment and combustion. The enrichment project was funded by N S F grant EAR 76-81598 Geochemistry Program. The critical comments of H Erlenkeuser (Kiel) substantially improved this paper.

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